

Subject:- Chemistry

Chemical Thermodynamics and Energetics

Practice Problems-II

Q1) Three moles of an ideal gas are expanded isothermally from a volume of 300 cm³ to 2.5 L at 300 K against a pressure of 1.9 atm. Calculate the work done in L atm and joules.

Soln:- $V_1 = 300 \text{ c.c.} = 0.3 \text{ L}$, $V_2 = 2.5 \text{ L}$, $T = 300 \text{ K}$, $P_{ex} = 1.9 \text{ atm}$

$$W = -P_{ex}\Delta V = -P_{ex}(V_2 - V_1) \\ = -1.9(2.5 - 0.3) = -1.9 \times 2.2$$

$$\therefore W = -4.18 \text{ L atm}$$

$$\text{and } W = -4.18 \times 101.3 = -423.4 \text{ J}$$

Q2) One mole of an ideal gas is compressed from 500 cm³ against a constant pressure of $1.216 \times 10^5 \text{ Pa}$. The work involved in the process is 36.50 J. Calculate the final volume.

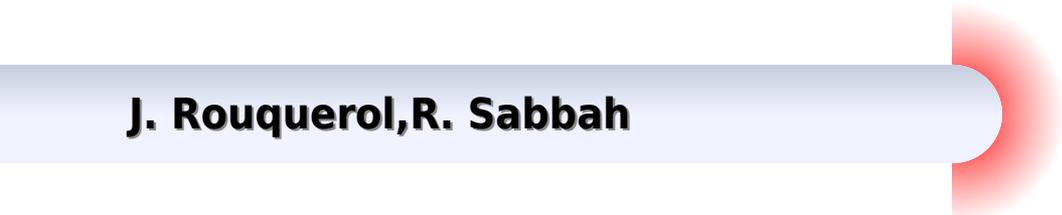
Soln:- $V_1 = 500 \text{ c.c.} = 0.5 \text{ dm}^3$,

$$P_{ex} = 1.216 \times 10^5 \text{ Pa} = 1.216 \times 10^2 \text{ kPa}, W = 36.50 \text{ J}$$

$$W = -P_{ex}\Delta V$$

Problems In Chemical Thermodynamics With Solutions

J. Rouquerol, R. Sabbah



Problems In Chemical Thermodynamics With Solutions:

Problems in Chemical Thermodynamics with Solutions Maka Aleksishvili, Shota Sidamonidze, 2002 The methods of chemical thermodynamics are effectively used in many fields of science and technology Mastering these methods and their use in practice requires profound comprehension of the theoretical questions and acquisition of certain calculating skills This book is useful to undergraduate and graduate students in chemistry as well as chemical thermal and refrigerating technology it will also benefit specialists in all other fields who are interested in using these powerful methods in their practical activities

Chemical Thermodynamics M L McGlashan, 2007-10-31 Specialist Periodical Reports provide systematic and detailed review coverage of progress in the major areas of chemical research Written by experts in their specialist fields the series creates a unique service for the active research chemist supplying regular critical in depth accounts of progress in particular areas of chemistry For over 80 years the Royal Society of Chemistry and its predecessor the Chemical Society have been publishing reports charting developments in chemistry which originally took the form of Annual Reports However by 1967 the whole spectrum of chemistry could no longer be contained within one volume and the series Specialist Periodical Reports was born The Annual Reports themselves still existed but were divided into two and subsequently three volumes covering Inorganic Organic and Physical Chemistry For more general coverage of the highlights in chemistry they remain a must Since that time the SPR series has altered according to the fluctuating degree of activity in various fields of chemistry Some titles have remained unchanged while others have altered their emphasis along with their titles some have been combined under a new name whereas others have had to be discontinued

Practical Chemical Thermodynamics for Geoscientists Bruce Fegley, Rose Osborne, 2013 Presents brief historical summaries and biographies of key thermodynamics scientists alongside the fundamentals they were responsible for

Basic Chemical Thermodynamics Eric Brian Smith, 2004 This widely acclaimed text now in its fifth edition and translated into many languages continues to present a clear simple and concise introduction to chemical thermodynamics An examination of equilibrium in the everyday world of mechanical objects provides the starting point for an accessible account of the factors that determine equilibrium in chemical systems This straightforward approach leads students to a thorough understanding of the basic principles of thermodynamics which are then applied to a wide range of physico chemical systems The book also discusses the problems of non ideal solutions and the concept of activity and provides an introduction to the molecular basis of thermodynamics Over five editions the views of teachers of the subject and their students have been incorporated The result is a little more rigour in specifying the dimensions within logarithmic expressions the addition of more worked examples and the inclusion of a simple treatment of the molecular basis of thermodynamics Students on courses in thermodynamics will continue to find this popular book an excellent introductory text

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biological molecules namely colossal dimension exclusive orderliness and capability to be in different states or conformations depending on conditions Organized into eight chapters this book begins with an overview of the trends in thermochemistry that involve complex reaction systems and product mixtures This text then discusses the problems relating to the standard state of solids and illustrates the utilization of enthalpy of mixing data Other chapters consider the available heat capacity results in the liquid gas This book discusses as well the high temperature measurement of thermodynamic data for substances of metallurgical interest The final chapter deals with the important advances in the experimental methods of heat capacity measurements including laser flash calorimetry and the high resolution heat capacity calorimeter This book is a valuable resource for chemists physical chemists thermochemists thermophysicists nuclear engineers and research workers

Bulletin of Chemical Thermodynamics, 1987 **Concepts And Problems In Physical Chemistry** P.S. Raghavan, 1997

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Edgeworth Fenwick, Thomas George Roddick, George Ross, 1906 Chemical Addresses Delivered at the Second Decennial Celebration of Clark University Clark University (Worcester, Mass.), 1911 Catalogue Johns Hopkins University. College for Teachers, 1930 *Catalog Issue* Pratt Institute. School of Science and Technology, 1926 *Science*, 1926 **An Introduction to the principles of physical chemistry from the standpoint of modern atomistics and thermodynamics** Edward Wight Washburn, 1915 *The Chemical News*, 1911 *Chemical News and Journal of Physical Science*, 1902 *Johns Hopkins University Circulars* Johns Hopkins University, 1928

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Table of Contents Problems In Chemical Thermodynamics With Solutions

1. Understanding the eBook Problems In Chemical Thermodynamics With Solutions
 - The Rise of Digital Reading Problems In Chemical Thermodynamics With Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Problems In Chemical Thermodynamics With Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Problems In Chemical Thermodynamics With Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Problems In Chemical Thermodynamics With Solutions
 - Personalized Recommendations
 - Problems In Chemical Thermodynamics With Solutions User Reviews and Ratings
 - Problems In Chemical Thermodynamics With Solutions and Bestseller Lists
5. Accessing Problems In Chemical Thermodynamics With Solutions Free and Paid eBooks
 - Problems In Chemical Thermodynamics With Solutions Public Domain eBooks
 - Problems In Chemical Thermodynamics With Solutions eBook Subscription Services
 - Problems In Chemical Thermodynamics With Solutions Budget-Friendly Options
6. Navigating Problems In Chemical Thermodynamics With Solutions eBook Formats

- ePub, PDF, MOBI, and More
 - Problems In Chemical Thermodynamics With Solutions Compatibility with Devices
 - Problems In Chemical Thermodynamics With Solutions Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Problems In Chemical Thermodynamics With Solutions
 - Highlighting and Note-Taking Problems In Chemical Thermodynamics With Solutions
 - Interactive Elements Problems In Chemical Thermodynamics With Solutions
 8. Staying Engaged with Problems In Chemical Thermodynamics With Solutions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Problems In Chemical Thermodynamics With Solutions
 9. Balancing eBooks and Physical Books Problems In Chemical Thermodynamics With Solutions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Problems In Chemical Thermodynamics With Solutions
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Problems In Chemical Thermodynamics With Solutions
 - Setting Reading Goals Problems In Chemical Thermodynamics With Solutions
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Problems In Chemical Thermodynamics With Solutions
 - Fact-Checking eBook Content of Problems In Chemical Thermodynamics With Solutions
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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