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web as a result of the reaction the equilibrium concentration of  $NH_3(g)$  became  $0.540\text{ M}$  what is the value of the equilibrium constant for this reaction at the given temperature of  $700\text{ C}$   $N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$  i write the initial concentrations of reactants and products

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