

STUDENT BOOK ANSWERS
To come
Chapter 9: Gas laws
Worked example 9.1

- 1 $n = \frac{V}{V_m} = \frac{2.33}{22.71} = 0.1026 \text{ mol}$
 $m = n \times MM = 0.1026 \times (12.01 + (4 \times 1.008)) = 1.646 = 1.65 \text{ g (3sf)}$
- 2 $n = \frac{m}{MM} = \frac{23.0}{14.01 + (3 \times 1.008)} = 1.350 \text{ mol}$
 $V = n \times V_m = 1.350 \times 24.79 = 33.47 = 33.5 \text{ L (3sf)}$

Worked example 9.2

- 1 $2\text{Na(s)} + 2\text{H}_2\text{O(l)} \rightarrow 2\text{NaOH(aq)} + \text{H}_2\text{(g)}$
 $n(\text{Na}) = \frac{m}{MM} = \frac{1.87}{22.99 + 16.00 + 1.008} = 0.04675 \text{ mol}$
 $n(\text{H}_2) = \frac{1}{2} \times n(\text{Na}) = 0.5 \times 0.04675 = 0.02338 \text{ mol}$
 $V = n \times V_m = 0.02338 \times 24.79 = 0.5795 = 0.580 \text{ L (3sf)}$
- 2 $n(\text{HCl}) = \frac{V}{V_m} = \frac{13.4}{22.71} = 0.5900 \text{ mol}$
 $2\text{KCl(aq)} + \text{H}_2\text{SO}_4\text{(aq)} \rightarrow 2\text{HCl(g)} + \text{K}_2\text{SO}_4\text{(aq)}$
 $n(\text{KCl}) = n(\text{HCl}) = 0.5900 \text{ mol}$
 $m(\text{KCl}) = n \times MM = 0.5900 \times (39.10 + 35.45) = 43.98 = 44.0 \text{ g (3sf)}$

Check your understanding 9.1/9.2

- 1 The volume of a gas changes with temperature and pressure, both of which compress or expand the gas particles.
- 2 Use the formula $n = \frac{V}{V_m}$ to convert the volume of gas to a number of moles. V_m depends upon the temperature of the gas. Use $m = n \times MM$ to convert the number of moles (n) to a mass (m) by multiplying by the molecular mass (MM).
- 3 **a** **i** $2 \times 250 = 500 \text{ mL}$
ii 250 mL
- b** **i** 500 mL
ii Nothing, as the law of combining volumes only applies to gases.

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Table of Contents Chapter 9 Notes Cengage Learning

1. Understanding the eBook Chapter 9 Notes Cengage Learning
 - The Rise of Digital Reading Chapter 9 Notes Cengage Learning
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 9 Notes Cengage Learning
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 9 Notes Cengage Learning
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 9 Notes Cengage Learning
 - Personalized Recommendations
 - Chapter 9 Notes Cengage Learning User Reviews and Ratings
 - Chapter 9 Notes Cengage Learning and Bestseller Lists

5. Accessing Chapter 9 Notes Cengage Learning Free and Paid eBooks
 - Chapter 9 Notes Cengage Learning Public Domain eBooks
 - Chapter 9 Notes Cengage Learning eBook Subscription Services
 - Chapter 9 Notes Cengage Learning Budget-Friendly Options
6. Navigating Chapter 9 Notes Cengage Learning eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 9 Notes Cengage Learning Compatibility with Devices
 - Chapter 9 Notes Cengage Learning Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 9 Notes Cengage Learning
 - Highlighting and Note-Taking Chapter 9 Notes Cengage Learning
 - Interactive Elements Chapter 9 Notes Cengage Learning
8. Staying Engaged with Chapter 9 Notes Cengage Learning
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 9 Notes Cengage Learning
9. Balancing eBooks and Physical Books Chapter 9 Notes Cengage Learning
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 9 Notes Cengage Learning
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 9 Notes Cengage Learning
 - Setting Reading Goals Chapter 9 Notes Cengage Learning
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 9 Notes Cengage Learning
 - Fact-Checking eBook Content of Chapter 9 Notes Cengage Learning
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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