

Flame Test Lab

Activity Key

Background

The heat from a laboratory burner will cause the ions of some elements to give off light. Electrons will absorb the heat energy from the flame and will “jump” to a higher energy level. When the electrons return to their original energy levels, this absorbed energy is released as light. Different elements absorb and release different amounts of energy and thus produce different kinds of light. This light may be visible or invisible, infrared or ultraviolet radiation. Analysis of this light can be used to identify elements and molecules. In today’s lab, we will observe the visible range of light emission.

Objectives

1. Perform a flame test to identify the characteristic color of metal ions.
2. Identify the metal ion in an unknown solution.
3. Calculate the energy of emitted photons.

Materials: (per lab group)

10 ml dilute solutions of the following:

Barium Chloride
Calcium Chloride
Lithium Chloride
Potassium Chloride
Sodium Chloride
Strontium Chloride
Unknown Solution

7 Wood Splints
1 Bunsen Burner
Safety Goggles & Aprons

Note:

If chloride compounds are not available, metal nitrate compounds may be substituted.

Use dilute or approximately 0.1 M solutions

Unknowns:

Number the beakers 1-12.
Add solutions as follows:

- | | |
|-----------------------|------------------------|
| 1. Strontium Chloride | 7. Calcium Chloride |
| 2. Calcium Chloride | 8. Strontium Chloride |
| 3. Potassium Chloride | 9. Barium Chloride |
| 4. Barium Chloride | 10. Potassium Chloride |
| 5. Lithium Chloride | 11. Sodium Chloride |
| 6. Sodium Chloride | 12. Lithium Chloride |

Procedure

1. Put on your safety goggles and an apron.
2. Obtain a wood splint and the barium chloride solution.
3. Dip the wood splint into the solution and allow it to soak for 1-2 minutes.
4. Light Bunsen Burner. USE CAUTION!
5. Remove the splint from the solution and hold the end of the splint into the flame. Do not burn the splint.
6. Observe and record the color of the flame.
7. Repeat steps 2-6 using the remaining chloride solutions.
8. Obtain an unknown solution from your teacher. Record the number of your unknown.
9. Soak a wood splint in the unknown solution.
10. Hold the splint in the flame and record the color of the flame that is produced.
11. Using your data, identify the metal ion in your unknown solution.

Missing Restaurant Lab Activity Answers

Lingsheng Yao



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Table of Contents Missing Restaurant Lab Activity Answers

1. Understanding the eBook Missing Restaurant Lab Activity Answers
 - The Rise of Digital Reading Missing Restaurant Lab Activity Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Missing Restaurant Lab Activity Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Missing Restaurant Lab Activity Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Missing Restaurant Lab Activity Answers
 - Personalized Recommendations

- Missing Restaurant Lab Activity Answers User Reviews and Ratings
- Missing Restaurant Lab Activity Answers and Bestseller Lists
- 5. Accessing Missing Restaurant Lab Activity Answers Free and Paid eBooks
 - Missing Restaurant Lab Activity Answers Public Domain eBooks
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 - Missing Restaurant Lab Activity Answers Budget-Friendly Options
- 6. Navigating Missing Restaurant Lab Activity Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Missing Restaurant Lab Activity Answers Compatibility with Devices
 - Missing Restaurant Lab Activity Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Missing Restaurant Lab Activity Answers
 - Highlighting and Note-Taking Missing Restaurant Lab Activity Answers
 - Interactive Elements Missing Restaurant Lab Activity Answers
- 8. Staying Engaged with Missing Restaurant Lab Activity Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Missing Restaurant Lab Activity Answers
- 9. Balancing eBooks and Physical Books Missing Restaurant Lab Activity Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Missing Restaurant Lab Activity Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Missing Restaurant Lab Activity Answers
 - Setting Reading Goals Missing Restaurant Lab Activity Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Missing Restaurant Lab Activity Answers
 - Fact-Checking eBook Content of Missing Restaurant Lab Activity Answers

- Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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