

Chemistry - Chapter 7 Study Guide

Section 7.1:

- *Valence electrons* – electrons in the outer energy level
 - responsible for bonding with other atoms
- How do you know how many valence electrons an element has?
 - Look at the group number
 - Group 1A = 1 valence electron
 - Group 7A = 7 valence electrons
 - Etc...
- *Electron Dot Diagrams* – a way of showing and keeping track of valence electrons
 - Written using the following steps:
 1. Write the element symbol – representing the nucleus and inner electrons
 2. Put one dot for each valence electrons,
 3. Don't forget to add them using Hund's rule – they don't pair up unless they have to.
- *The Octet Rule* – atoms tend to want to achieve a noble gas configuration because it is the most stable. Octet = eight electrons
- *Cations* – a positive ion (usually metals)
 - Lose electrons to attain a noble gas configuration
 - Written like Na^{1+}
 - No dots are shown (on an electron dot diagram) because they lose the valence electrons they had
- *Anions* – a negative ion (usually non-metals)
 - Gain electrons to attain a noble gas configuration
 - Written like P^{3-}
 - Dots are shown because valence electrons are gained
- 8 valence electrons = stable atom (noble gases.)

Section 7.2:

- Anions (-) and cations (+) are held together by opposite charges; a bond is formed through the transfer of electrons to achieve a noble gas configuration (octet.)
- *Ionic Compounds:*
 - Also called salts
 - Made from a cation and an anion
 - Which means they're made of a metal and a non-metal
- ***if you don't understand how to do ionic bonding, check the book or ask somebody to help... it's hard to explain***
- The result of ionic bonding is an ionic compound, shown in a *formula unit*
 - Written like Ca_3P_2

Chemistry Matter And Change Chapter 7 Study Guide Answer Key

R Sandford



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Table of Contents Chemistry Matter And Change Chapter 7 Study Guide Answer Key

1. Understanding the eBook Chemistry Matter And Change Chapter 7 Study Guide Answer Key
 - The Rise of Digital Reading Chemistry Matter And Change Chapter 7 Study Guide Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Matter And Change Chapter 7 Study Guide Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Matter And Change Chapter 7 Study Guide Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Matter And Change Chapter 7 Study Guide Answer Key
 - Personalized Recommendations
 - Chemistry Matter And Change Chapter 7 Study Guide Answer Key User Reviews and Ratings
 - Chemistry Matter And Change Chapter 7 Study Guide Answer Key and Bestseller Lists

5. Accessing Chemistry Matter And Change Chapter 7 Study Guide Answer Key Free and Paid eBooks
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 - Highlighting and Note-Taking Chemistry Matter And Change Chapter 7 Study Guide Answer Key
 - Interactive Elements Chemistry Matter And Change Chapter 7 Study Guide Answer Key
8. Staying Engaged with Chemistry Matter And Change Chapter 7 Study Guide Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Matter And Change Chapter 7 Study Guide Answer Key
9. Balancing eBooks and Physical Books Chemistry Matter And Change Chapter 7 Study Guide Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Matter And Change Chapter 7 Study Guide Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Matter And Change Chapter 7 Study Guide Answer Key
 - Setting Reading Goals Chemistry Matter And Change Chapter 7 Study Guide Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Matter And Change Chapter 7 Study Guide Answer Key
 - Fact-Checking eBook Content of Chemistry Matter And Change Chapter 7 Study Guide Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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