

Dimensional Analysis Practice Worksheet Solutions

Complete this and place it in your portfolio after section 2-2. I will email the solutions later so that you can check your work (make sure you are showing the work the same way you see it in the solutions, using dimensional analysis). It is important that you practice and learn to use dimensional analysis (multiplying by unit fractions) appropriately. Make sure you get help if you are still having difficulties with this topic. Use my office hours or the math tutors in the Student Learning Center in the library. You do not have to write sentences on this practice worksheet. You do need to show all of the units and use \approx for approximations.

Do the following conversions using dimensional analysis using the conversion factors on page 74 of your text. You should be doing these all in one computation (even if you have to use more than one unit fraction), not several computations. I have done #1 as example.

1. 5400 inches to miles

$$5400 \cancel{\text{in.}} \times \frac{1 \cancel{\text{ft}}}{12 \cancel{\text{in.}}} \times \frac{1 \text{ mi}}{5280 \cancel{\text{ft}}} = \frac{5400}{63360} \text{ mi} \\ \approx 0.085 \text{ mi}$$

2. 13 weeks to seconds

$$13 \cancel{\text{ weeks}} \times \frac{7 \cancel{\text{ days}}}{1 \cancel{\text{ week}}} \times \frac{24 \cancel{\text{ hrs}}}{1 \cancel{\text{ day}}} \times \frac{60 \cancel{\text{ min}}}{1 \cancel{\text{ hr}}} \times \frac{60 \text{ sec}}{1 \cancel{\text{ min}}} = 7,862,400 \text{ sec}$$

3. 54 yards to mm

$$54 \cancel{\text{ yards}} \times \frac{3 \cancel{\text{ feet}}}{1 \cancel{\text{ yd}}} \times \frac{1 \cancel{\text{ m}}}{3.28 \cancel{\text{ ft}}} \times \frac{1000 \text{ mm}}{1 \cancel{\text{ m}}} = \frac{162000}{3.28} \text{ mm} \\ \approx 49,390.24 \text{ mm}$$

4. $36 \frac{\text{cm}}{\text{sec}}$ to mph

$$\frac{36 \cancel{\text{ cm}}}{1 \cancel{\text{ sec}}} \times \frac{60 \cancel{\text{ sec}}}{1 \text{ hr}} \times \frac{1 \cancel{\text{ m}}}{1000 \cancel{\text{ cm}}} \times \frac{3.28 \cancel{\text{ ft}}}{1 \cancel{\text{ m}}} \times \frac{1 \text{ mi}}{5280 \cancel{\text{ ft}}} = \frac{7084.8 \text{ mi}}{5280000 \text{ hr}} \\ \approx 0.00134 \frac{\text{mi}}{\text{hr}}$$

Chemistry Dimensional Analysis Practice Problems Answers

**John W. Moore, Conrad L.
Stanitski, Peter C. Jurs**



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of chemical concepts as well as the right mix of general chemistry organic chemistry and biochemistry An emphasis on real world topics lets readers clearly see how the chemistry will apply to their career

Dosage and Solutions Calculations Alice J. Smith,1989

General Chemistry James P. Birk,Donald M. Baird,1994-03

Chemistry, Student Solutions Manual James E. Brady,Joel W. Russell,John R. Holum,2000-01-26

The images on the cover call attention to the relationship between macro observations and the intimate structure of chemical substances and the changes both chemical and physical that they undergo

Fireworks One of the ingredients is phosphorus a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms

The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light

Carbon The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball shaped molecules often referred to as buckeyballs

Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth

Acetylene torch Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen a reaction between molecules of C_2H_2 and O_2 to give CO_2 and H_2O

Hot air balloon The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and eight hydrogen atoms

Stormy weather The evaporation of water serves to store energy provided by the sun

Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet

General Chemistry, Reactions First Kevin Revell,2024-12-04

Revell s General Chemistry empowers students to grasp essential topics and concepts with more ease

Using a friendly approach the text uses metaphors and relatable examples to demystify even the most challenging subjects in general chemistry

Preparatory Chemistry Howard Stephen Stoker,1990

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Understanding Chemistry Fred M. Dewey,1994

Selected Titles in Chemistry ,1972

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Problem Solving in General Chemistry Ronald A. DeLorenzo,1981

Chemistry John W. Moore,Conrad L. Stanitski,Peter C. Jurs,2002

CHEMISTRY THE MOLECULAR SCIENCE is intended to help students develop a broad overview of chemistry and chemical reactions an understanding of the most important concepts and models that chemists and those in chemistry related fields use an appreciation of the many ways chemistry impacts our daily lives the ability to apply the facts concepts and models of chemistry appropriately to new situations in chemistry other sciences and engineering and to other disciplines

Chemistry Theodore Lawrence Brown,1988

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