

## Chapter 9 Practice Test - Naming and Writing Chemical Formulas

### Matching

Match each item with the correct statement below.

Match each item with the correct statement below.

- |                                |                    |
|--------------------------------|--------------------|
| a. monatomic ion               | f. cation          |
| b. acid                        | g. binary compound |
| c. base                        | h. anion           |
| d. law of definite proportions | i. polyatomic ion  |
| e. law of multiple proportions |                    |

- \_\_\_ 1. consists of a single atom with a positive or negative charge
- \_\_\_ 2. atom or group of atoms having a negative charge
- \_\_\_ 3. atom or group of atoms having a positive charge
- \_\_\_ 4. tightly-bound group of atoms that behaves as a unit and carries a net charge
- \_\_\_ 5. compound composed of two different elements
- \_\_\_ 6. produces a hydrogen ion when dissolved in water
- \_\_\_ 7. produces a hydroxide ion when dissolved in water
- \_\_\_ 8. In any chemical compound, the masses of elements are always in the same proportion by mass.
- \_\_\_ 9. when two elements form more than one compound, the masses of one element that combine with the same mass of the other element are in the ratio of small, whole numbers

### Multiple Choice

Identify the choice that best completes the statement or answers the question.

- \_\_\_ 10. When naming a transition metal ion that can have more than one common ionic charge, the numerical value of the charge is indicated by a \_\_\_\_.
- |           |                                     |
|-----------|-------------------------------------|
| a. prefix | c. Roman numeral following the name |
| b. suffix | d. superscript after the name       |
- \_\_\_ 11. Which of the following correctly provides the name of the element, the symbol for the ion, and the name of the ion?
- |                                   |                                     |
|-----------------------------------|-------------------------------------|
| a. fluorine, $F^+$ , fluoride ion | c. copper, $Cu^+$ , cuprous ion     |
| b. zinc, $Zn^{2+}$ , zincate ion  | d. sulfur, $S^{2-}$ , sulfurous ion |
- \_\_\_ 12. The nonmetals in Groups 6A and 7A \_\_\_\_.
- lose electrons when they form ions
  - have a numerical charge that is found by subtracting 8 from the group number
  - all have ions with a  $-1$  charge
  - end in *-ite*
- \_\_\_ 13. Which of the following is NOT a cation?
- |                  |                  |
|------------------|------------------|
| a. iron(III) ion | c. $Ca^{2+}$     |
| b. sulfate       | d. mercurous ion |

# Chapter 9 Practice Test Naming And Writing Chemical Formulas

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