

Covalent Bonding Worksheet

Covalent bonding occurs when two or more NON_METALS share electrons, attempting to attain a stable octet (8 outer electrons) in their outer shell for at least part of the time. Draw a Lewis dot diagram for each element listed. Circle the unpaired electrons that will be shared between the elements.

1.) H ₂ hydrogen is diatomic: $\text{H} + \text{H} \longrightarrow \text{H} \cdot \cdot \text{H} \longrightarrow \text{H}-\text{H} \quad \text{Single Bond}$
2.) F ₂ fluorine is diatomic: $\text{F} + \text{F} \longrightarrow \cdot \cdot \text{F} \cdot \cdot \cdot \cdot \text{F} \cdot \cdot \longrightarrow$
3.) O ₂ oxygen is diatomic: $\text{O} + \text{O} \longrightarrow \quad \quad \quad \longrightarrow \text{O}=\text{O} \quad \text{Double Bond}$
4.) N ₂ nitrogen is diatomic. Is this a triple bond? $\text{N} + \text{N}$
5.) BF ₃ you need 3 fluorine atoms here: $\text{B} + \text{F}$
6.) Ammonia NH ₃ hint: how many hydrogen atoms are needed? $\text{N} + \text{H}$
7.) Carbon dioxide CO ₂ $\text{O} + \text{C} + \text{O}$
8.) Methane CH ₄ careful here 4 hydrogen atoms needed $\text{C} + \text{H}$
9.) Dihydrogen monoxide: the most dangerous substance on the planet. It has killed more people than any other substance known to mankind!! $\text{H} + \text{O} + \text{H}$
10.) SO ₂ hint: one pair of electrons from sulfur must be slit up for this one to work. $\text{O} + \text{S} + \text{O}$

Chapter 8 Covalent Bonding Worksheet Answers

R Barnett



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