



COVALENT BONDING

Practice Problems

Do your answers to the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either an ionic or a molecular compound.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - C_2H_6
 - SO_2
 - CaF_2
 - CH_4

- Classify each of the following as either an ionic or a molecular compound.
 - K_2SO_4
 - C_2H_6
 - H_2O
 - HCl
 - CaF_2
 - SO_2
 - CH_4
 - C_2H_6
- Write the formula for each of the following molecular compounds.
 - Carbon dioxide
 - Carbon monoxide
 - Diphosphorus pentoxide
 - Diphosphorus trioxide
 - Diphosphorus pentasulfide
 - Diphosphorus trisulfide
 - Diphosphorus tetrasulfide
 - Diphosphorus hexasulfide
 - Diphosphorus heptasulfide
 - Diphosphorus octasulfide
 - Diphosphorus nonasulfide
 - Diphosphorus decasulfide

SECTION 8.2 THE NATURE OF COVALENT BONDING

- Show covalent bonding in hydrogen fluoride. $\text{H}-\text{F}$
- Show covalent bonding in methane. C with four H atoms bonded to it.
- Show covalent bonding in carbon dioxide. $\text{O}=\text{C}=\text{O}$
- Show covalent bonding in acetylene. $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw Lewis structures for each of the following.
 - H_2O
 - CO_2
 - CH_4
 - C_2H_6
 - C_2H_2
 - C_2H_4
 - $\text{C}_2\text{H}_2\text{Cl}_2$
 - $\text{C}_2\text{H}_2\text{Cl}_4$
 - $\text{C}_2\text{H}_2\text{Cl}_6$
 - $\text{C}_2\text{H}_2\text{Cl}_8$
 - $\text{C}_2\text{H}_2\text{Cl}_{10}$
 - $\text{C}_2\text{H}_2\text{Cl}_{12}$
 - $\text{C}_2\text{H}_2\text{Cl}_{14}$
 - $\text{C}_2\text{H}_2\text{Cl}_{16}$
 - $\text{C}_2\text{H}_2\text{Cl}_{18}$
 - $\text{C}_2\text{H}_2\text{Cl}_{20}$
 - $\text{C}_2\text{H}_2\text{Cl}_{22}$
 - $\text{C}_2\text{H}_2\text{Cl}_{24}$
 - $\text{C}_2\text{H}_2\text{Cl}_{26}$
 - $\text{C}_2\text{H}_2\text{Cl}_{28}$
 - $\text{C}_2\text{H}_2\text{Cl}_{30}$
 - $\text{C}_2\text{H}_2\text{Cl}_{32}$
 - $\text{C}_2\text{H}_2\text{Cl}_{34}$
 - $\text{C}_2\text{H}_2\text{Cl}_{36}$
 - $\text{C}_2\text{H}_2\text{Cl}_{38}$
 - $\text{C}_2\text{H}_2\text{Cl}_{40}$
 - $\text{C}_2\text{H}_2\text{Cl}_{42}$
 - $\text{C}_2\text{H}_2\text{Cl}_{44}$
 - $\text{C}_2\text{H}_2\text{Cl}_{46}$
 - $\text{C}_2\text{H}_2\text{Cl}_{48}$
 - $\text{C}_2\text{H}_2\text{Cl}_{50}$
 - $\text{C}_2\text{H}_2\text{Cl}_{52}$
 - $\text{C}_2\text{H}_2\text{Cl}_{54}$
 - $\text{C}_2\text{H}_2\text{Cl}_{56}$
 - $\text{C}_2\text{H}_2\text{Cl}_{58}$
 - $\text{C}_2\text{H}_2\text{Cl}_{60}$
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 - $\text{C}_2\text{H}_2\text{Cl}_{76}$
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 - $\text{C}_2\text{H}_2\text{Cl}_{92}$
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 - $\text{C}_2\text{H}_2\text{Cl}_{96}$
 - $\text{C}_2\text{H}_2\text{Cl}_{98}$
 - $\text{C}_2\text{H}_2\text{Cl}_{100}$

SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and suborbital are represented by each of the following?
 - $2s$
 - $2p_x$
 - $2p_y$
 - $2p_z$
 - $3s$
 - $3p_x$
 - $3p_y$
 - $3p_z$
 - $4s$
 - $4p_x$
 - $4p_y$
 - $4p_z$
 - $5s$
 - $5p_x$
 - $5p_y$
 - $5p_z$
- Which orbital and suborbital are represented by each of the following?
 - $1s$
 - $2s$
 - $2p_x$
 - $2p_y$
 - $2p_z$
 - $3s$
 - $3p_x$
 - $3p_y$
 - $3p_z$
 - $4s$
 - $4p_x$
 - $4p_y$
 - $4p_z$
 - $5s$
 - $5p_x$
 - $5p_y$
 - $5p_z$

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Karin Nielsen-Saines



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