

**Stoichiometry Worksheet 2**  
**Chemistry**

Name \_\_\_\_\_

Date \_\_\_\_\_ Hour \_\_\_\_\_

1. In the reaction :  $2\text{Al}_2\text{O}_{3(s)} \rightarrow 4\text{Al}_{(s)} + 3\text{O}_{2(g)}$ , the mole ratios are:

$\text{Al}_2\text{O}_3$  to Al \_\_\_\_\_  $\text{Al}_2\text{O}_3$  to  $\text{O}_2$  \_\_\_\_\_ Al to  $\text{O}_2$  \_\_\_\_\_  $\text{O}_2$  to Al \_\_\_\_\_

2. In the reaction:  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ , the mole ratio of oxygen to water is: \_\_\_\_\_

3. In the reaction:  $\text{Ca} + \text{Cl}_2 \rightarrow \text{CaCl}_2$ , the mole ratio of chlorine to calcium is: \_\_\_\_\_

4. In the reaction:  $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2$ , the mole ratio of zinc to sulfuric acid is:  
\_\_\_\_\_

5. The equation for the Haber process is represented by:



To convert 9.0 moles of hydrogen to ammonia would require how many moles of nitrogen?

6. In the equation:  $2\text{KClO}_{3(g)} \rightarrow 2\text{KCl}_{(s)} + 3\text{O}_{2(g)}$ , how many moles of oxygen are produced when 3.00 moles of  $\text{KClO}_3$  decompose completely?

7. For the reaction:  $\text{C}_{(s)} + 2\text{H}_{2(g)} \rightarrow \text{CH}_{4(g)}$ , how many moles of  $\text{H}_2$  are required to produce 10.0 moles of  $\text{CH}_4$ ?

## Chapter 3 Stoichiometry Exercises Answers

**M Mark**



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