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1

Chemistry Lab Experiment

Determination of the Molar Volume of a Gas

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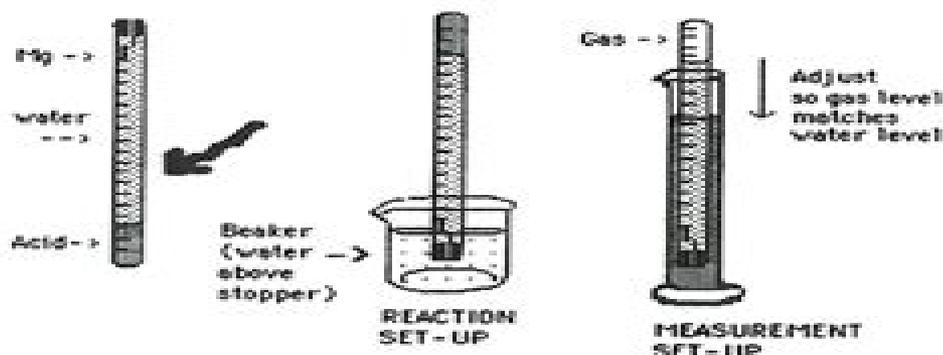
Name:

Lab Partner(s):

Purpose:

- To experimentally determine the molar volume of hydrogen gas and to compare it with the $\frac{22.4\text{-L}}{\text{mole}}$ theoretical STP value.
- To relate gas law concepts learned in MA Frameworks Topic 6 (e.g., the Combined Gas Law and Dalton's Law of Partial Pressures) to previously learned concepts involving balancing chemical equations and stoichiometric relationships.

Introduction:



The **water displacement method** can be used to capture the hydrogen gas liberated when a strip of magnesium metal reacts with a 3.0-*M* solution of hydrochloric acid. The total gas volume of H₂O vapor and H₂ adjusted to the room's barometric pressure can be determined from the **eudiometer** (graduated gas-measuring tube) by partially immersing the tube into a large container of water until the liquid level inside the tube matches the water level in the large container.

According to **Dalton's Law of Partial Pressures**, the total gas pressure in the eudiometer is the sum of the pressure of the hydrogen gas collected plus the **water vapor pressure**. Symbolically, $P_{\text{total}} = P_{\text{H}_2} + P_{\text{H}_2\text{O}}$. Subtracting the water vapor pressure from the total pressure gives the pressure of the dry hydrogen gas.

Determining Molar Volume Gas Post Lab Answers

**James M. Postma, Julian L. Robert, J.
Leland Hollenberg**



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Working with Chemistry Donald J. Wink, Sharon Fetzer-Gislason, Julie Ellefson Kuehn, 2004-02-20 With this modular laboratory program students build skills using important chemical concepts and techniques to the point where they are able to design a solution to a scenario drawn from a professional environment The scenarios are drawn from the lives of people who work with chemistry every day ranging from field ecologists to chemical engineers and include many health professionals as well

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