

$$\text{Mg} \rightarrow \text{Mg}^{2+} + 2e^-$$

element	# of protons	# of electrons	ending electron distribution	# of valence e	family #	to become stable	# of protons	# of electrons	ion charge
Li	<u>3</u> p ⁺	<u>3</u> e ⁻	2s ¹	1	1	(lose) <u>1</u> e ⁻	___ p ⁺	___ e ⁻	
Be	<u>4</u> p ⁺	<u>4</u> e ⁻	2s ²	2	2	(lose) <u>2</u> e ⁻	___ p ⁺	___ e ⁻	
B	<u>5</u> p ⁺	<u>5</u> e ⁻	2p ¹	3	13	(lose) <u>3</u> e ⁻	___ p ⁺	___ e ⁻	
C	<u>6</u> p ⁺	<u>6</u> e ⁻	2p ²	4	14	(lose) <u>4</u> e ⁻	___ p ⁺	___ e ⁻	
N	<u>7</u> p ⁺	<u>7</u> e ⁻	2p ³	5	15	(lose) <u>5</u> e ⁻	___ p ⁺	___ e ⁻	

15

[illegible]

CHEMISTRY: A Study of Matter

10

Worksheet Predicting Ionic Charges Answers

JS Bruner



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