

## AP Chemistry Chapter 4 Practice Exam

D

1. The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.

How much sulfur (in grams) is produced from 21.5 grams of  $\text{O}_2$ ?

A) 97.0 g

B) 14.4 g

C) 21.5 g

D) 43.1 g

E) none of these

$$21.5 \text{ g O}_2 \left| \frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \right| \frac{2 \text{ SO}_2}{3 \text{ O}_2} \left| \frac{3 \text{ S}}{\text{SO}_2} \right| \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 43.1 \text{ g S}$$

B

2. Iron is produced from its ore by the reactions:

How many moles of  $\text{O}_2(g)$  are needed to produce 4.6 moles of  $\text{Fe}(s)$ ?A) 2.3 mole  $\text{O}_2$ B) 3.5 mole  $\text{O}_2$ C) 4.6 mole  $\text{O}_2$ D) 6.9 mole  $\text{O}_2$ E) 13.8 mole  $\text{O}_2$ 

$$4.6 \text{ mol Fe}(s) \left| \frac{3 \text{ CO}}{2 \text{ Fe}} \right| \frac{1 \text{ O}_2}{2 \text{ CO}} = 3.45 \text{ mol O}_2$$

B

3. Suppose the reaction
- $\text{Ca}_3(\text{PO}_4)_2 + 3\text{H}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{H}_3\text{PO}_4$
- is carried out starting with 153 g of
- $\text{Ca}_3(\text{PO}_4)_2$

and 76.8 g of  $\text{H}_2\text{SO}_4$ . How much phosphoric acid will be produced?

A) 76.7 g

B) 31.1 g

C) 229.8 g

D) 115.1 g

E) 96.7 g

$$153 \text{ g Ca}_3\text{PO}_4)_2 \left| \frac{1 \text{ mol Ca}_3\text{PO}_4)_2}{310.18 \text{ g}} \right| \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ Ca}_3\text{PO}_4)_2} =$$

$$76.8 \text{ g H}_2\text{SO}_4 \left| \frac{1 \text{ mol H}_2\text{SO}_4}{98.08 \text{ g}} \right| \frac{2 \text{ H}_3\text{PO}_4}{3 \text{ H}_2\text{SO}_4} =$$

B

- 4.
- $\text{SO}_2$
- reacts with
- $\text{H}_2\text{S}$
- as follows:

When 7.50 g of  $\text{H}_2\text{S}$  reacts with 12.75 g of  $\text{SO}_2$ , which statement applies?

A) 6.38 g of sulfur are formed.

B) 10.6 g of sulfur are formed.

C) 0.0216 moles of  $\text{H}_2\text{S}$  remain.D) 1.13 g of  $\text{H}_2\text{S}$  remain.E)  $\text{SO}_2$  is the limiting reagent.

$$7.50 \text{ g H}_2\text{S} \left| \frac{1 \text{ mol H}_2\text{S}}{34.03 \text{ g H}_2\text{S}} \right| \frac{3 \text{ S}}{2 \text{ H}_2\text{S}} \left| \frac{32 \text{ g S}}{1 \text{ mol S}} \right| = 10.89$$

$$12.75 \text{ g SO}_2 \left| \frac{1 \text{ mol SO}_2}{64.06 \text{ g SO}_2} \right| \frac{3 \text{ S}}{1 \text{ SO}_2} \left| \frac{32 \text{ g S}}{1 \text{ mol S}} \right| = 19.13$$

# Ap Chemistry Test Answers

**M Mark**



## **Ap Chemistry Test Answers:**

This book delves into Ap Chemistry Test Answers. Ap Chemistry Test Answers is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Ap Chemistry Test Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
  - Chapter 1: Introduction to Ap Chemistry Test Answers
  - Chapter 2: Essential Elements of Ap Chemistry Test Answers
  - Chapter 3: Ap Chemistry Test Answers in Everyday Life
  - Chapter 4: Ap Chemistry Test Answers in Specific Contexts
  - Chapter 5: Conclusion
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3. In chapter 2, this book will delve into the foundational concepts of Ap Chemistry Test Answers. The second chapter will elucidate the essential principles that must be understood to grasp Ap Chemistry Test Answers in its entirety.
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6. In chapter 5, this book will draw a conclusion about Ap Chemistry Test Answers. This chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Ap Chemistry Test Answers.

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