

# CHEMISTRY



**Properties of Matter**

- Matter is anything that has mass and occupies space.
- Matter is made up of particles.
- Particles are very small.
- Particles are continuously moving.
- Particles attract each other.

**States of Matter**

- Solid: Particles are closely packed and have a fixed shape and volume.
- Liquid: Particles are less closely packed and have a fixed volume but no fixed shape.
- Gas: Particles are far apart and have neither a fixed shape nor a fixed volume.

**Changes of State**

- Melting: Solid to liquid.
- Freezing: Liquid to solid.
- Boiling: Liquid to gas.
- Condensation: Gas to liquid.
- Sublimation: Solid to gas.
- Deposition: Gas to solid.

**Atomic Structure**

Atoms are the smallest particles that cannot be created or destroyed in a chemical reaction. They are made up of three sub-particles: electrons, protons, and neutrons.

**Electrons**

- Discovered by Cathode ray experiment.
- Negatively charged.
- Present in all atoms.

**Protons**

- Discovered by Rutherford's alpha particle scattering experiment.
- Positively charged.
- Present in all atoms.

**Neutrons**

- Discovered by Chadwick.
- No charge.
- Present in all atoms except hydrogen.

**Atomic Structure (Continued)**

Atoms are electrically neutral. The number of protons is equal to the number of electrons. The mass of an atom is concentrated in the nucleus, which contains protons and neutrons.

**Atomic Models**

- Thomson's model: Plum pudding model.
- Rutherford's model: Nuclear model.
- Bohr's model: Planetary model.

Element	Atomic Number (Z)	Mass Number (A)
Hydrogen	1	1
Helium	2	4
Lithium	3	7
Boron	5	11
Carbon	6	12
Nitrogen	7	14
Oxygen	8	16
Fluorine	9	19
Neon	10	20

**Atomic Structure (Continued)**

Electrons are arranged in shells or energy levels around the nucleus. The number of electrons in each shell is limited. The outermost shell is called the valence shell.

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**Sebastian Brünink**



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