

# Solubility

## Introduction

The purpose of this lab is to explore the relationship between temperature and solubility using a laboratory procedure. What is the effect of temperature on the solubility of a solid in a liquid? If the temperature of the liquid is increased, then more sugar will dissolve because warm solutions hold more solute than cold solutions. The independent variable for this lab is the temperature of the water and the dependent variable is the amount of sugar that can be dissolved.

## Materials and Procedures

The materials needed for this lab are:

1. 100 mL beaker
2. beaker tongs
3. spoon
4. stirring rod
5. thermometer
6. hot plate
7. hot pad
8. balance
9. 300 g sugar
10. ice
11. water

The procedures for this lab are as follows:

1. Measure the Mass of One Spoonful of Sugar
  - a. Use a weighing paper to measure the number of grams of sugar that are in one spoonful of sugar. Be sure to use a level spoonful, not a heaping spoonful, to help ensure that each time you add sugar, it is approximately the same amount. Record that value above the data table.
2. Measure Solubility in Cold Water
  - a. Place ice in a 100 mL beaker. Fill the beaker with water until the ice and water total 50 mL. Stir until just a small amount of ice remains. Let the temperature of the chilled water stabilize to obtain a temperature = 5°C (between 0°C and 10°C).
  - b. Record the temperature in °C.
  - c. Add a spoonful of sugar (again, a level spoonful, not a heaping spoonful). Stir using the stirring rod. Continue adding spoonfuls of sugar until the solution becomes saturated, stirring after each spoonful is added.
  - d. After the solution reaches saturation, record the amount of sugar (in spoonfuls) dissolved before the saturation point. Do not discard the solution; continue to use it throughout the rest of the procedure.

# Solubility Lab Experiment With Answers

**BM King**

A decorative graphic element consisting of a light blue horizontal bar with a rounded right end, and a red-to-white gradient semi-circle overlapping the right side of the bar.

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