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A Solution Manual for:  
COMBUSTION, 4th Ed.

by

I. Glassman, and R.A. Yetter

CHAPTER 1

**Problem 1**

Calculate the heat of reaction when methane and air in stoichiometric proportions are brought into a calorimeter at 500 K. The product composition is brought to the ambient temperature (298 K) by the cooling water. The pressure in the calorimeter is assumed to remain at 1 atm, but the water formed has condensed.

**Solution:** The heat of reaction can be calculated using Eqn. (10) in the text. However, as also suggested in the textbook, this equation can be simplified for the general case when  $H_{T_2}^o = H_{T_{amb}}^o$ . The modified equation becomes:

$$-Q = \Delta H = \sum_{i \in \text{prod}} n_i \left[ (H_{T_2}^o - H_{T_{amb}}^o) + (\Delta H_f^o)_{T_{amb}} \right] - \sum_{i \in \text{react}} n_i \left[ (H_{T_1}^o - H_{T_{amb}}^o) + (\Delta H_f^o)_{T_{amb}} \right] \quad (1)$$

In this Equation,  $T_2$  and  $T_1$  are the temperatures of the products and reactants respectively. The general equation for the complete combustion of one mole of methane in air is:



where  $a_{\text{air}}$  is a coefficient describing how much air is needed for stoichiometric combustion. Balancing the above reaction one has:



In order to apply Eqn. (1) to the above reaction, the following properties from Tables in the textbook are needed (note that for this problem  $T_2 = 298 \text{ K}$  and  $T_1 = 500 \text{ K}$ ):

gases	$\Delta H_f^o$ (kJ/mol)	$H_{T_1}^o - H_{T_{amb}}^o$ (kJ/mol)	$H_{T_2}^o - H_{T_{amb}}^o$ (kJ/mol)
CH <sub>4</sub>	-74.90	8.21	0.0
O <sub>2</sub>	0.0	6.09	0.0
N <sub>2</sub>	0.0	5.92	0.0
H <sub>2</sub> O <sub>(liq)</sub>	-286.04	0.0	0.0
CO <sub>2</sub>	-393.77	0.0	0.0

Equation (1) then becomes:

2

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# Glassman Combustion Solution Manual

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