

Geomorphology

- Geomorphology is the study of the physical processes and surface features of the Earth

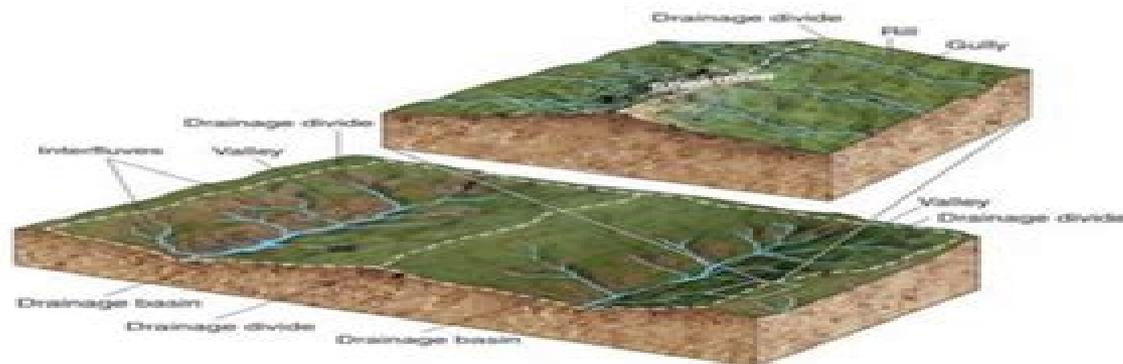
What happens to the precipitation that falls within a drainage basin?

- The drainage basin is an open system, of inputs, stores, flows and output
- The drainage basin collects precipitation and allows water to seep into the ground through the process of infiltration to become underground water. water is stored in lakes in drainage basins.

Features of a drainage basin

- Catchment area
- River system
- Tributary
- Watershed
- River mouth
- Source
- Confluence
- interfluvium
- water table
- Ground water
- Surface run-off
- infiltration

(NB: Check the definitions in the glossary on the first page)



Geomorphology A Level Notes

M. Schwartz



Geomorphology A Level Notes :

Reviewing **Geomorphology A Level Notes** : Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Geomorphology A Level Notes** ," an enthralling opus penned by a very acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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 $2 \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + \text{C}_2\text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g}) + 2 \text{H}_2\text{O}(\text{g})$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{g})$ $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g})$ $\Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$ $\Delta h = 1411 \text{ kJ}$ mole
 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l})$ $\Delta h = 1560 \text{ kJ}$ mole
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta h = 285 \text{ kJ}$ mole

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 hess s law example 12.1 determine Δh for the target reaction $\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$ given the following information

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 $\text{Mg}(\text{s}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{MgO}(\text{s})$ we used

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 cohen 2016 the equations for the reactions in the experiment done are as follows
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web thus for reaction 9 2 in which 1 mole of h_3o and 1 mole of oh combine to form 2 moles of h_2o $\Delta H = 55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{products} - \Delta H_{reactants}$ is positive if heat is absorbed that is if $\Delta H_{products}$

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web hess s law enthalpy of reaction and hess s law abstract the goal of this lab is to perform experiments and practice hess s law by determining enthalpy change in chemical reactions in this lab three experiments were conducted in order to test hess s law the experiments required observing and collecting data on three different reactions

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions 1 $\text{NaOH (s)} \rightarrow \text{Na}^+ \text{(aq)} + \text{OH}^- \text{(aq)}$ 2 $\text{NaOH (s)} + \text{H}^+ \text{(aq)} \rightarrow \text{Na}^+ \text{(aq)} + \text{H}_2\text{O (l)}$ 3