

Acids, Bases, Buffers and pH Data Tables

Part A.

| | Solution | pH Meter | | pH paper | | Universal Indicator | |
|----|--------------------|----------|--------------------------------------|------------------|----------|---------------------|----------|
| | | Reading | $[H_3O^+]$ | color | pH value | color | pH value |
| | Control (H_2O) | | | | | | |
| a. | 0.1 M NH_3 | 11.00 | $10^{-11} = 1^{-11}$ | blue | 12 | purple | 10 |
| b. | 0.1 M NaOH | 12.85 | $10^{-12.85} = 1.41 \times 10^{-13}$ | blue | 13 | purple | 10 |
| c. | 0.1 M H_3BO_3 | 6.5 | $10^{-6.5} = 3.16 \times 10^{-7}$ | orange yellow | 4.5 | yellow | 6.5 |
| d. | 0.1 M HCl | 1.40 | $10^{-1.40} = .039$ | orange | 3 | red/pink | 3 |

Part B.

| | Solution | pH Meter | | Universal Indicator | |
|----|----------------------|----------|--------------------------------------|---------------------|----------|
| | | Reading | $[H_3O^+]$ | color | pH value |
| | Control (H_2O) | | | | |
| a. | 0.1 M Na_2CO_3 | 11.42 | $10^{-11.42} = 3.80 \times 10^{-12}$ | purple | 10 |
| b. | 0.1 M $(NH_4)_2SO_4$ | 6.58 | $10^{-6.58} = 2.63 \times 10^{-7}$ | yellow | 6 |
| c. | 0.1 M NaCl | 6.80 | $10^{-6.80} = 1.58 \times 10^{-7}$ | bright green | 6.5 |

Part C.

| | Solution | pH Meter |
|----|--------------------|----------|
| | | Reading |
| | Control (H_2O) | |
| a. | Acetic Acid | 3.05 |
| b. | Sodium Acetate | 7.15 |
| c. | Mixture of a & b | 4.83 |

Ph Properties Of Buffer Solutions Pre Lab Answers

David Kirk



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