

## CHAPTER 15 REVIEW

## Chemical Equilibrium

## Teacher Notes and Answers

Question 10  
Equilibrium 1

## Answer: A and B

$$K_c = \frac{[\text{CO}_2][\text{H}_2]}{[\text{CO}][\text{H}_2\text{O}]}$$

11. a. The value of the forward equilibrium constant,  $K_c$ , is related to the reverse value,  $K_c$ , by the relationship:

$K_c = 1/K_c$  (where  $K_c$  is the forward equilibrium constant and  $K_c$  is the reverse equilibrium constant).

$$K_c = 1/K_c$$

$$K_c = 1/K_c$$

## Question 12

## Answer: A and B

$$K_c = \frac{[\text{CO}_2][\text{H}_2]}{[\text{CO}][\text{H}_2\text{O}]}$$

13. These results are also obtained by considering the equilibrium constant for the reverse reaction,  $K_c = 1/K_c$ .

14. The value of  $K_c$  is independent of the concentration of the reactants and products.

15. a. The value of  $K_c$  is independent of the concentration of the reactants and products. The value of  $K_c$  is independent of the concentration of the reactants and products.

b. The value of  $K_c$  is independent of the concentration of the reactants and products. The value of  $K_c$  is independent of the concentration of the reactants and products.

c. The value of  $K_c$  is independent of the concentration of the reactants and products.

16. The value of  $K_c$  is independent of the concentration of the reactants and products. The value of  $K_c$  is independent of the concentration of the reactants and products.

17. The value of  $K_c$  is independent of the concentration of the reactants and products. The value of  $K_c$  is independent of the concentration of the reactants and products.

## Question 13

## Answer: A and B

$$K_c = \frac{[\text{CO}_2][\text{H}_2]}{[\text{CO}][\text{H}_2\text{O}]}$$

18. The value of  $K_c$  is independent of the concentration of the reactants and products.

$$K_c = \frac{[\text{CO}_2][\text{H}_2]}{[\text{CO}][\text{H}_2\text{O}]}$$

# Chapter 18 Review Chemical Equilibrium Section 3 Answers

**Jiyuan Zhang**



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