

## Chapter 17 Thermochemistry Study Guide

### 17.1 – 17.2 Thermochemical Equations

- Make the following conversions:
  - 444 cal to joules =  $1.86 \times 10^3 \text{ J}$
  - 1.8 kJ to joules =  $1.8 \times 10^3 \text{ J}$
  - 0.45 kJ to calories =  $1.1 \times 10^2 \text{ cal}$
- Classify each of these processes as endothermic or exothermic:
  - condensing steam – **exo**
  - burning alcohol – **exo**
  - evaporating alcohol – **endo**
  - baking a potato – **endo**
- The specific heat capacity for silver is  $0.24 \text{ J/g}^\circ\text{C}$ . Calculate the energy required to raise the temperature of 150.0 g Ag from 273 K to 298 K. Calculate the molar heat capacity of silver.

$$\text{Energy: } q = (150.0)(.24)(298-273) = 9.0 \times 10^2 \text{ J}$$

$$\text{Molar heat capacity: } (0.24 \text{ J/g}^\circ\text{C})(107.87 \text{ g/mol Ag}) = 26 \text{ J/mol}$$

- It takes 585 J of energy to raise the temperature of 125.6 g Hg from  $20.0^\circ\text{C}$  to  $53.5^\circ\text{C}$ . Calculate the specific heat capacity and the molar heat capacity of Hg.

$$\text{Specific heat capacity: } C = q/m\Delta t = 585/(125.6)(53.5-20) = 0.139 \text{ J/g}^\circ\text{C}$$

$$\text{Molar heat capacity: } (0.139 \text{ J/g}^\circ\text{C})(200.59 \text{ g/mol Hg}) = 27.9 \text{ J/mol}$$

- A 46.2-g sample of copper is heated to  $95.4^\circ\text{C}$  and then placed in a calorimeter containing 75.0 g water at  $19.6^\circ\text{C}$ . The equilibrium temperature in the calorimeter is  $21.8^\circ\text{C}$ . Calculate the specific heat capacity of copper; assuming that all the heat lost by the copper is gained by the water.

$$C = 0.203 \text{ J/g}^\circ\text{C}$$

- A 15.0-g sample of nickel metal is heated to  $100.0^\circ\text{C}$  and dropped into 55.0 g of water, initially at  $23.0^\circ\text{C}$ . Assuming that all the heat lost by the nickel is absorbed by the water; calculate the final temperature of the nickel and the water. The specific heat of nickel is  $0.444 \text{ J/g}^\circ\text{C}$ .

$$\begin{array}{rcl} 15.0(.444)(100.0 - T_f) & = & 55.0(4.184)(T_f - 23.0) \\ 666 - 6.66T_f & = & 230.12T_f - 5292.76 \\ 5958.76 & = & 236.78 T_f \\ & \mathbf{25.2^\circ\text{C}} & = T_f \end{array}$$

# Chapter 17 Thermochemistry Study Guide Answers

**Emilie Sanchez**



## **Chapter 17 Thermochemistry Study Guide Answers:**

## Decoding **Chapter 17 Thermochemistry Study Guide Answers**: Revealing the Captivating Potential of Verbal Expression

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