

17

THERMOCHEMISTRY

Reaction that releases heat is exothermic
 Reaction that absorbs heat is endothermic
 Exothermic: $\Delta H < 0$
 Endothermic: $\Delta H > 0$

Practice Problems

Use the following information for the following problems.

SECTION 17.1 THE FLOW OF ENERGY—HEAT AND WORK

Use the following information for the following problems.

1. How much energy is released when 1.00 mol of C_2H_6 is burned in excess oxygen?
2. How much energy is released when 1.00 mol of C_2H_6 is burned in excess oxygen? (Assume $\Delta H_{\text{comb}}^\circ(\text{C}_2\text{H}_6) = -84.7 \text{ kJ/mol}$)
3. How much energy is released when 1.00 mol of C_2H_6 is burned in excess oxygen? (Assume $\Delta H_{\text{comb}}^\circ(\text{C}_2\text{H}_6) = -84.7 \text{ kJ/mol}$)



SECTION 17.2 MEASURING AND EXPRESSING ENTHALPY CHANGES

1. Calculate the enthalpy change for the reaction of 1.00 mol of C_2H_6 with 1.00 mol of O_2 to form $\text{C}_2\text{H}_6\text{O}$ and H_2O . (Assume $\Delta H_{\text{comb}}^\circ(\text{C}_2\text{H}_6) = -84.7 \text{ kJ/mol}$)
2. Calculate the enthalpy change for the reaction of 1.00 mol of C_2H_6 with 1.00 mol of O_2 to form $\text{C}_2\text{H}_6\text{O}$ and H_2O . (Assume $\Delta H_{\text{comb}}^\circ(\text{C}_2\text{H}_6) = -84.7 \text{ kJ/mol}$)
3. Calculate the enthalpy change for the reaction of 1.00 mol of C_2H_6 with 1.00 mol of O_2 to form $\text{C}_2\text{H}_6\text{O}$ and H_2O . (Assume $\Delta H_{\text{comb}}^\circ(\text{C}_2\text{H}_6) = -84.7 \text{ kJ/mol}$)
4. Calculate the enthalpy change for the reaction of 1.00 mol of C_2H_6 with 1.00 mol of O_2 to form $\text{C}_2\text{H}_6\text{O}$ and H_2O . (Assume $\Delta H_{\text{comb}}^\circ(\text{C}_2\text{H}_6) = -84.7 \text{ kJ/mol}$)

SECTION 17.3 HEAT IN CHANGES OF STATE

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M Mark



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