

## Chapter 12

## Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- \_\_\_\_\_ 1. The first step in most stoichiometry problems is to \_\_\_\_\_.  
a. add the coefficients of the reagents  
b. convert given quantities to moles  
c. convert given quantities to volumes  
d. convert given quantities to masses
- \_\_\_\_\_ 2. In the reaction  $2\text{CO}(g) + \text{O}_2(g) \rightarrow 2\text{CO}_2(g)$ , what is the ratio of moles of oxygen used to moles of  $\text{CO}_2$  produced?  
a. 1:1  
b. 2:1  
c. 1:2  
d. 2:2
- \_\_\_\_\_ 3. How many moles of glucose,  $\text{C}_6\text{H}_{12}\text{O}_6$ , can be "burned" biologically when 10.0 mol of oxygen is available?  
 $\text{C}_6\text{H}_{12}\text{O}_6(s) + 6\text{O}_2(g) \rightarrow 6\text{CO}_2(g) + 6\text{H}_2\text{O}(l)$   
a. 0.938 mol  
b. 1.67 mol  
c. 53.3 mol  
d. 60.0 mol
- \_\_\_\_\_ 4. At STP, how many liters of oxygen are required to react completely with 3.6 liters of hydrogen to form water?  
 $2\text{H}_2(g) + \text{O}_2(g) \rightarrow 2\text{H}_2\text{O}(g)$   
a. 1.8 L  
b. 3.6 L  
c. 2.0 L  
d. 2.4 L
- \_\_\_\_\_ 5. The equation below shows the decomposition of lead nitrate. How many grams of oxygen are produced when 11.5 g  $\text{NO}_2$  is formed?  
 $2\text{Pb}(\text{NO}_3)_2(s) \rightarrow 2\text{PbO}(s) + 4\text{NO}_2(g) + \text{O}_2(g)$   
a. 1.00 g  
b. 2.00 g  
c. 2.88 g  
d. 32.0 g
- \_\_\_\_\_ 6. Iron(III) oxide is formed when iron combines with oxygen in the air. How many grams of  $\text{Fe}_2\text{O}_3$  are formed when 16.7 g of Fe reacts completely with oxygen?  
 $4\text{Fe}(s) + 3\text{O}_2(g) \rightarrow 2\text{Fe}_2\text{O}_3(s)$   
a. 12.0 g  
b. 23.9 g  
c. 47.8 g  
d. 95.6 g
- \_\_\_\_\_ 7. When glucose is consumed, it reacts with oxygen in the body to produce carbon dioxide, water, and energy. How many grams of carbon dioxide would be produced if 45 g of  $\text{C}_6\text{H}_{12}\text{O}_6$  completely reacted with oxygen?  
a. 1.5 g  
b. 1.8 g  
c. 11 g  
d. 66 g
- \_\_\_\_\_ 8. Aluminum reacts with sulfuric acid to produce aluminum sulfate and hydrogen gas. How many grams of aluminum sulfate would be formed if 250 g  $\text{H}_2\text{SO}_4$  completely reacted with aluminum?  
 $2\text{Al}(s) + 3\text{H}_2\text{SO}_4(aq) \rightarrow \text{Al}_2(\text{SO}_4)_3(aq) + 3\text{H}_2(g)$   
a. 0.85 g  
b. 290 g  
c. 450 g  
d. 870 g

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**M Mark**



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