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Chapter 17 & 19 Study Guide

Key Terms

Directions: Define each of the following key terms from chapters seventeen and nineteen on a separate sheet of paper.

- **Atom** - the smallest piece of matter that still has the properties of the element
- **atomic number** – number of protons in an atom's nucleus
- **atomic mass** – average mass of all an element's isotopes
- **ductile** – capable of being drawn into wire
- **electron** – subatomic particle with a negative charge
- **electron cloud** – area outside the nucleus in which electrons can be found.
- **Electron dot diagram** – uses the element's symbol and dots to represent the valence electrons in an element.
- **energy level** – areas within the electron cloud where electrons are most likely to be found.
- **Family** – vertical columns of the periodic table
- **Group** - vertical columns of the periodic table
- **half-life** – the amount of time it takes half the sample of a radioactive isotope to decay.
- **Isotope** – atoms of the same element that have different numbers of neutrons.
- **Lustrous** – having a shiny appearance
- **Malleable** – can be hammered flat or rolled into sheets
- **mass number** – sum of protons and neutrons in the nucleus of an atom.
- **Metal** – element that typically is hard, shiny, solid, malleable, and is a good conductor of heat and electricity.
- **Neutron** – neutrally charged subatomic particle.
- **Nonmetal** - element that typically is brittle solid or gas, not shiny, not malleable, and is a poor conductor of heat and electricity.
- **Nucleus** – center of the atom
- **Period** – horizontal row of the periodic table.
- **periodic table** – organized list of all known elements.
- **Proton** – positively charged subatomic particle
- **transition metal** – Elements found in families 3-12 of the periodic table.
- **transuranium element** – an element containing more than 92 protons in its nucleus
- **quark** – particle believed to make up protons and neutrons.
- **valence electron** – electrons found in the outermost energy level of an atom.

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Christian G. Meyer



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Table of Contents Chapter 17 Study Guide Answers

1. Understanding the eBook Chapter 17 Study Guide Answers
 - The Rise of Digital Reading Chapter 17 Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 17 Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 17 Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 17 Study Guide Answers
 - Personalized Recommendations
 - Chapter 17 Study Guide Answers User Reviews and Ratings
 - Chapter 17 Study Guide Answers and Bestseller Lists

5. Accessing Chapter 17 Study Guide Answers Free and Paid eBooks
 - Chapter 17 Study Guide Answers Public Domain eBooks
 - Chapter 17 Study Guide Answers eBook Subscription Services
 - Chapter 17 Study Guide Answers Budget-Friendly Options
6. Navigating Chapter 17 Study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 17 Study Guide Answers Compatibility with Devices
 - Chapter 17 Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 17 Study Guide Answers
 - Highlighting and Note-Taking Chapter 17 Study Guide Answers
 - Interactive Elements Chapter 17 Study Guide Answers
8. Staying Engaged with Chapter 17 Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 17 Study Guide Answers
9. Balancing eBooks and Physical Books Chapter 17 Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 17 Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 17 Study Guide Answers
 - Setting Reading Goals Chapter 17 Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 17 Study Guide Answers
 - Fact-Checking eBook Content of Chapter 17 Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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