

Chapter 17: Chemical Equilibrium

Worksheet: K<sub>c</sub> and K<sub>p</sub>

1. Consider the reaction:  $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$

Problem 1

At 1000 K, the equilibrium constant  $K_c$  for the reaction is 0.25.

- a. Write the expression for  $K_c$  for the reaction.
- b. Calculate the equilibrium concentration of  $\text{SO}_3$  if the initial concentrations are  $[\text{SO}_2] = 0.40 \text{ M}$  and  $[\text{O}_2] = 0.20 \text{ M}$ .
- c. Calculate the equilibrium concentration of  $\text{SO}_2$  if the initial concentrations are  $[\text{SO}_2] = 0.40 \text{ M}$  and  $[\text{O}_2] = 0.20 \text{ M}$ .
- d. Calculate the equilibrium concentration of  $\text{O}_2$  if the initial concentrations are  $[\text{SO}_2] = 0.40 \text{ M}$  and  $[\text{O}_2] = 0.20 \text{ M}$ .
- e. Calculate the equilibrium concentration of  $\text{SO}_3$  if the initial concentrations are  $[\text{SO}_2] = 0.40 \text{ M}$  and  $[\text{O}_2] = 0.20 \text{ M}$ .
- f. Calculate the equilibrium concentration of  $\text{SO}_2$  if the initial concentrations are  $[\text{SO}_2] = 0.40 \text{ M}$  and  $[\text{O}_2] = 0.20 \text{ M}$ .
- g. Calculate the equilibrium concentration of  $\text{O}_2$  if the initial concentrations are  $[\text{SO}_2] = 0.40 \text{ M}$  and  $[\text{O}_2] = 0.20 \text{ M}$ .
- h. Calculate the equilibrium concentration of  $\text{SO}_3$  if the initial concentrations are  $[\text{SO}_2] = 0.40 \text{ M}$  and  $[\text{O}_2] = 0.20 \text{ M}$ .
- i. Calculate the equilibrium concentration of  $\text{SO}_2$  if the initial concentrations are  $[\text{SO}_2] = 0.40 \text{ M}$  and  $[\text{O}_2] = 0.20 \text{ M}$ .
- j. Calculate the equilibrium concentration of  $\text{O}_2$  if the initial concentrations are  $[\text{SO}_2] = 0.40 \text{ M}$  and  $[\text{O}_2] = 0.20 \text{ M}$ .

Problem 2

- 1. The equilibrium constant  $K_c$  for the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$  is 0.50 at 400 K. Calculate the equilibrium concentration of  $\text{NH}_3$  if the initial concentrations are  $[\text{N}_2] = 0.10 \text{ M}$  and  $[\text{H}_2] = 0.30 \text{ M}$ .
- 2. The equilibrium constant  $K_c$  for the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$  is 0.50 at 400 K. Calculate the equilibrium concentration of  $\text{N}_2$  if the initial concentrations are  $[\text{N}_2] = 0.10 \text{ M}$  and  $[\text{H}_2] = 0.30 \text{ M}$ .
- 3. The equilibrium constant  $K_c$  for the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$  is 0.50 at 400 K. Calculate the equilibrium concentration of  $\text{H}_2$  if the initial concentrations are  $[\text{N}_2] = 0.10 \text{ M}$  and  $[\text{H}_2] = 0.30 \text{ M}$ .
- 4. The equilibrium constant  $K_c$  for the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$  is 0.50 at 400 K. Calculate the equilibrium concentration of  $\text{NH}_3$  if the initial concentrations are  $[\text{N}_2] = 0.10 \text{ M}$  and  $[\text{H}_2] = 0.30 \text{ M}$ .
- 5. The equilibrium constant  $K_c$  for the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$  is 0.50 at 400 K. Calculate the equilibrium concentration of  $\text{N}_2$  if the initial concentrations are  $[\text{N}_2] = 0.10 \text{ M}$  and  $[\text{H}_2] = 0.30 \text{ M}$ .
- 6. The equilibrium constant  $K_c$  for the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$  is 0.50 at 400 K. Calculate the equilibrium concentration of  $\text{H}_2$  if the initial concentrations are  $[\text{N}_2] = 0.10 \text{ M}$  and  $[\text{H}_2] = 0.30 \text{ M}$ .
- 7. The equilibrium constant  $K_c$  for the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$  is 0.50 at 400 K. Calculate the equilibrium concentration of  $\text{NH}_3$  if the initial concentrations are  $[\text{N}_2] = 0.10 \text{ M}$  and  $[\text{H}_2] = 0.30 \text{ M}$ .
- 8. The equilibrium constant  $K_c$  for the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$  is 0.50 at 400 K. Calculate the equilibrium concentration of  $\text{N}_2$  if the initial concentrations are  $[\text{N}_2] = 0.10 \text{ M}$  and  $[\text{H}_2] = 0.30 \text{ M}$ .
- 9. The equilibrium constant  $K_c$  for the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$  is 0.50 at 400 K. Calculate the equilibrium concentration of  $\text{H}_2$  if the initial concentrations are  $[\text{N}_2] = 0.10 \text{ M}$  and  $[\text{H}_2] = 0.30 \text{ M}$ .
- 10. The equilibrium constant  $K_c$  for the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$  is 0.50 at 400 K. Calculate the equilibrium concentration of  $\text{NH}_3$  if the initial concentrations are  $[\text{N}_2] = 0.10 \text{ M}$  and  $[\text{H}_2] = 0.30 \text{ M}$ .

# Elements Of Language First Course Answer Key

**JA Banks**



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