

1. A certain gas obeys the equation of state  $U(S, V, N) = aS^7/V^4N^2$ , where  $a$  is a dimensioned constant. Here  $U$  represents the internal energy of the gas,  $S$  the entropy,  $V$  the volume and  $N$  the fixed number of particles of the system.

- (a) Let such a gas be filled in a box of volume  $V$  and the internal energy of the system be  $U$ . A partition is placed to divide the box into two equal parts, each having volume  $V/2$ . For each part, the internal energy is now  $\alpha U$  and the dimensioned constant be  $\beta a$ . Obtain  $\alpha$  and  $\beta$ . [3]

$\alpha =$	$\beta =$
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- (b) The temperature  $T$  can be expressed in terms of the derivative of internal energy as [2]

$$T = \left( \frac{dU}{dS} \right)_{V,N}$$

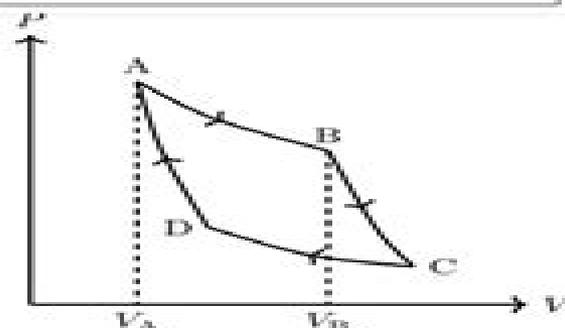
where the subscripts indicate that the differentiation has been carried out keeping  $V$  and  $N$  constant. In a similar way, express pressure  $P$  in terms of a derivative of the internal energy.

$P =$
-------

- (c) Find the equation of state of the given system relating  $P$ ,  $T$ , and  $V$ . [1]

$P =$
-------

- (d) One mole of this gas executes a Carnot cycle ABCDA between reservoirs at temperatures  $T_1$  and  $T_2$  ( $T_1 > T_2$ ). Obtain the heat change in the process AB ( $Q_{AB}$ ) and work done by the system in the processes AB and BC ( $W_{AB}$ ,  $W_{BC}$ ) of the cycle. Express your answers only in terms of temperatures  $T_1$ ,  $T_2$ , volumes  $V_A$ ,  $V_B$ , and the other constants. [7]



$Q_{AB} =$
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