



# Chapter Test Answers

## CHAPTER 1: CHEMICAL REACTIONS AND EQUATIONS

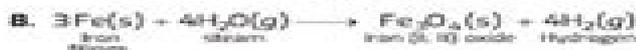
- (c) 4 and 2
- (d) double displacement reaction
- (a) only (i)
- (b)  $H_2$ , CuO
- (a) Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of Assertion (A).
- (c) Assertion (A) is true but Reason (R) is false.

7. (i) When zinc metal is added to a solution of copper sulphate, a chemical reaction occurs where zinc displaces copper from the copper sulphate solution, forming zinc sulphate and solid copper metal. As a result, the blue colour of the copper sulphate solution fades, and a reddish-brown deposit of copper metal forms on the zinc.

(ii) When magnesium ribbon reacts with dilute hydrochloric acid, magnesium chloride salt will be formed and hydrogen gas is liberated. The reaction is exothermic, so the solution becomes warmer.

(iii) Yes, a displacement reaction is also a redox reaction because it involves the transfer of electrons where one substance is oxidised and another is reduced.

(iv) Galvanisation, electroplating and extraction of metals are some applications of displacement reactions in everyday life or industry.



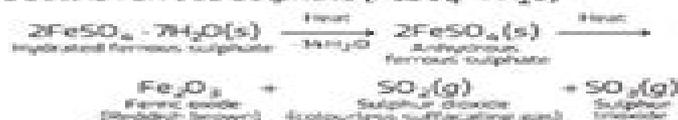
9. The given reaction is a redox reaction.

Oxidising agent =  $Fe_2O_3$

Reducing agent = Al

- Decomposition reaction
- Displacement reaction
- Double displacement reaction
- Combination Reaction

11. The salt is ferrous sulphate ( $FeSO_4 \cdot 7H_2O$ )



12. Exothermic reaction is the reaction in which heat is released.

Endothermic reaction is the reaction in which heat is absorbed.

(i) Heating of coal in air to form carbon dioxide → Exothermic reaction.

(ii) Heating of limestone in a lime kiln to form quick lime → Endothermic reaction.

13. (i) The reaction will take place in presence of sunlight.

This reaction is used in black and white photography.

(ii) (a) Displacement reaction because  $Cl_2$  is displacing  $I_2$  from KI solution.

(b) Combination reaction because K combines with  $Cl_2$  to form a single product, KCl.

14.

	Reaction	Substance	
		Oxidised	reduced
(i)	$\text{ZnO}_{(s)} + \text{C}_{(s)} \xrightarrow[\text{reduced}]{\text{Oxidised}} \text{Zn}_{(s)} + \text{CO}_{(g)}$	C	ZnO
(ii)	$4\text{Na}_{(s)} + \text{O}_{2(g)} \xrightarrow[\text{reduced}]{\text{Oxidised}} 2\text{Na}_2\text{O}_{(s)}$	Na	$O_2$
(iii)	$\text{CuO}_{(s)} + \text{H}_2(g) \xrightarrow[\text{Oxidised}]{\text{reduced}} \text{Cu}_{(s)} + \text{H}_2\text{O}(l)$	$H_2$	CuO

15. (i) X-copper (Cu). Process - Corrosion

(ii) Green coating - Basic copper carbonate [ $CuCO_3 \cdot Cu(OH)_2$ ]

(iii) Painting and oiling.

16. (i) **Decomposition reaction:** Reaction in which a compound splits up into two or more simpler substances.

(ii) **Oxidation reaction:** The reaction in which oxygen is added to a substance or hydrogen is removed from a substance.

(iii) **Displacement reaction:** Reaction in which a more reactive element displaces a less reactive element from its compound.

(iv) **Displacement reaction:** Reaction in which a more reactive element displaces a less reactive element from its compound.

(v) **Combination reaction:** Reaction in which two or more substances combine to form a single substance.

17. The process in which new substances with new properties are formed by the rearrangement of atoms is known as chemical reaction.

Four observations which helps us to determine that a chemical reaction has taken place are as follows:

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