

## Periodic Trends Lab

Periodic law states that when the elements are arranged in order of increasing atomic number, there is a periodic repetition of their chemical and physical properties. Elements within the same group have the most similar chemical and physical properties, although not identical. Elements in neighboring groups may also have similar chemical and physical properties, but to a lesser degree. Trends in properties occur within groups and across periods.

**PURPOSE:** In this activity, you will

- Interpret group and periodic trends in atomic radius, electronegativity, and ionization energy.
- Graph these trends.

**PROCEDURE:**

1. Complete the pre-lab section.
2. Using the data from **Table S** of your reference tables, prepare the graphs listed below. Staple your graphs to the back of this lab packet.

*On each graph do your best:*

- USE PENCIL!
- Put an appropriate title and labels for each axis.
- Omit the transition elements.
- Connect the points you plot.
- Connect the points representing Group 1 elements with a green line.
- Connect the points representing Group 2 elements with a blue line.
- Connect the points representing Group 17 elements with a red line.
- Draw a triangle around the points representing the elements of Period 3.

GRAPH 1. Atomic radius vs. atomic number for elements 1-54.

GRAPH 2. Electronegativity vs. atomic number for elements 1-54.

GRAPH 3. Ionization energy vs. atomic number for elements 1-54.

3. Complete the conclusion section.

# Periodic Trends Lab Question Answers

**Marcel A. Müller**



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## **Table of Contents Periodic Trends Lab Question Answers**

1. Understanding the eBook Periodic Trends Lab Question Answers
  - The Rise of Digital Reading Periodic Trends Lab Question Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Lab Question Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Trends Lab Question Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Lab Question Answers
  - Personalized Recommendations
  - Periodic Trends Lab Question Answers User Reviews and Ratings
  - Periodic Trends Lab Question Answers and Bestseller Lists
5. Accessing Periodic Trends Lab Question Answers Free and Paid eBooks
  - Periodic Trends Lab Question Answers Public Domain eBooks
  - Periodic Trends Lab Question Answers eBook Subscription Services
  - Periodic Trends Lab Question Answers Budget-Friendly Options

6. Navigating Periodic Trends Lab Question Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Periodic Trends Lab Question Answers Compatibility with Devices
  - Periodic Trends Lab Question Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Trends Lab Question Answers
  - Highlighting and Note-Taking Periodic Trends Lab Question Answers
  - Interactive Elements Periodic Trends Lab Question Answers
8. Staying Engaged with Periodic Trends Lab Question Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Trends Lab Question Answers
9. Balancing eBooks and Physical Books Periodic Trends Lab Question Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodic Trends Lab Question Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Periodic Trends Lab Question Answers
  - Setting Reading Goals Periodic Trends Lab Question Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Trends Lab Question Answers
  - Fact-Checking eBook Content of Periodic Trends Lab Question Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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