

STUDENT BOOK ANSWERS
To come
Chapter 9: Gas laws
Worked example 9.1

- 1 $n = \frac{V}{V_m} = \frac{2.33}{22.71} = 0.1026 \text{ mol}$
 $m = n \times MM = 0.1026 \times (12.01 + (4 \times 1.008)) = 1.646 = 1.65 \text{ g (3sf)}$
- 2 $n = \frac{m}{MM} = \frac{23.0}{14.01 + (3 \times 1.008)} = 1.350 \text{ mol}$
 $V = n \times V_m = 1.350 \times 24.79 = 33.47 = 33.5 \text{ L (3sf)}$

Worked example 9.2

- 1 $2\text{Na(s)} + 2\text{H}_2\text{O(l)} \rightarrow 2\text{NaOH(aq)} + \text{H}_2\text{(g)}$
 $n(\text{Na}) = \frac{m}{MM} = \frac{1.87}{22.99 + 16.00 + 1.008} = 0.04675 \text{ mol}$
 $n(\text{H}_2) = \frac{1}{2} \times n(\text{Na}) = 0.5 \times 0.04675 = 0.02338 \text{ mol}$
 $V = n \times V_m = 0.02338 \times 24.79 = 0.5795 = 0.580 \text{ L (3sf)}$
- 2 $n(\text{HCl}) = \frac{V}{V_m} = \frac{13.4}{22.71} = 0.5900 \text{ mol}$
 $2\text{KCl(aq)} + \text{H}_2\text{SO}_4\text{(aq)} \rightarrow 2\text{HCl(g)} + \text{K}_2\text{SO}_4\text{(aq)}$
 $n(\text{KCl}) = n(\text{HCl}) = 0.5900 \text{ mol}$
 $m(\text{KCl}) = n \times MM = 0.5900 \times (39.10 + 35.45) = 43.98 = 44.0 \text{ g (3sf)}$

Check your understanding 9.1/9.2

- 1 The volume of a gas changes with temperature and pressure, both of which compress or expand the gas particles.
- 2 Use the formula $n = \frac{V}{V_m}$ to convert the volume of gas to a number of moles. V_m depends upon the temperature of the gas. Use $m = n \times MM$ to convert the number of moles (n) to a mass (m) by multiplying by the molecular mass (MM).
- 3 **a** **i** $2 \times 250 = 500 \text{ mL}$
ii 250 mL
b **i** 500 mL
ii Nothing, as the law of combining volumes only applies to gases.

Chapter 9 Notes Cengage Learning

J Dewey



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Governance in Family Enterprises A. Koeberle-Schmid, D. Kenyon-Rouvinez, E. Poza, INOBUCHI, 2013-12-02 Presents a comprehensive overview of governance in family enterprises including practical management knowledge in easy to use frameworks and interviews with renowned family enterprise owners and managers Readers will benefit from the book's systematic approach and the opportunity to learn from the experience of other family enterprises 1, 2, 3

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