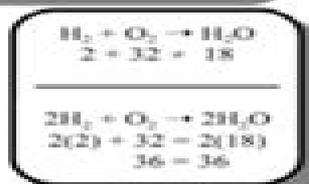


Test Review No 7

Conservation of Mass. Matter is neither created nor destroyed. During a chemical reaction the mass does not change. A properly written equation shows conservation of mass. For example, $\text{AgNO}_3(\text{aq}) + \text{NaCl}(\text{aq}) \rightarrow \text{NaNO}_3(\text{aq}) + \text{AgCl}(\text{s})$.

<p><u>Silver Nitrate</u> AgNO_3 Ag = $1 \times 108 = 108$ N = $1 \times 14 = 14$ O = $3 \times 16 = 48$ 170</p>	<p><u>Sodium Chloride</u> NaCl Na = $1 \times 23 = 23$ Cl = $1 \times 35 = 35$ 58</p>	<p><u>Sodium Nitrate</u> NaNO_3 Na = $1 \times 23 = 23$ N = $1 \times 14 = 14$ O = $3 \times 16 = 48$ 85</p>	<p><u>Silver Chloride</u> AgCl Ag = $1 \times 108 = 108$ Cl = $1 \times 35 = 35$ 143</p>
<p>$\text{AgNO}_3(\text{aq})$ + 170 +</p>	<p>$\text{NaCl}(\text{aq})$ → 58 =</p>	<p>$\text{NaNO}_3(\text{aq})$ + 85 +</p>	<p>$\text{AgCl}(\text{s})$ 143</p>
	228	228	228

Balancing Equations. The equation at the top of the box to the right does *not* show conservation of mass. Starting with two molecules of hydrogen, as shown in the equation at the bottom of the box by writing a coefficient 2 in front of the hydrogen and forming two molecules of water by writing a coefficient 2 in front of the water shows conservation. Coefficients are used to **balance** equations. Coefficients make the number of atoms of each type the same on the reactant and product side. As a result, coefficients make the mass the same on the reactant and product side of the equation. Balancing is done by counting the number and type of atoms on the reactant and product side of the equation and making them equal.



Moles. A mole is a formula mass expressed in grams. (1 mole = 1 gram formula mass). Atomic mass units are too small to measure on a laboratory balance, but grams are not. An atom of carbon has a mass of 12 amu and a molecule of glucose has a mass of 180 amu. Each mass represents one particle. Since the mass ratios in formula masses and gram formula masses are the same (12 amu:180 amu::12 g:180 g), the ratio of particles must still be the same (1mole:1 mole). The gram formula mass (GFM) is the number of grams in 1 mole. This results in the mathematical relationships shown above and to the right.

Substance	Formula Mass	Gram Formula Mass
carbon	12 amu	12 g
sodium chloride (NaCl)	58 amu	58 g
glucose (C ₆ H ₁₂ O ₆)	180 amu	180 g

- $GFM = \frac{g}{mole}$
- $g = GFM \times mole$
- $mole = \frac{g}{GFM}$

Stoichiometry. Stoichiometry is the branch of chemistry that deals with the application of the laws of definite proportions and of the conservation of mass and energy to chemical activity. It shows the quantitative relationship between constituents of a chemical reaction. Stoichiometric calculations are based on several assumptions. It is assumed that the reaction has no side reactions, the reaction goes to completion, and the reactants are completely consumed. One type of problem that can be solved stoichiometrically is based on the mole ratios of a balanced equation. A sample problem is shown to the right.

Sample Problem

How many moles of oxygen are consumed when 0.6 moles of hydrogen burns to produce water?

Step 1: Write a balanced equation and determine the mole ratios from the equation

$$2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}$$

mole ratio	2	1	2
moles	known 0.6	unknown x	

Step 2: Identify the known and the unknown

Step 3: Set up a proportion and solve for the unknown

$$\propto \frac{2}{0.6\text{mol}} = \frac{1}{x}$$

$$\propto 2x = 0.6\text{mol}$$

$$\propto x = 0.3\text{mol}$$

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JR Anderson



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Table of Contents Evan P Silberstein 2003 Answer Key

1. Understanding the eBook Evan P Silberstein 2003 Answer Key
 - The Rise of Digital Reading Evan P Silberstein 2003 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Evan P Silberstein 2003 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Evan P Silberstein 2003 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Evan P Silberstein 2003 Answer Key
 - Personalized Recommendations

- Evan P Silberstein 2003 Answer Key User Reviews and Ratings
- Evan P Silberstein 2003 Answer Key and Bestseller Lists
- 5. Accessing Evan P Silberstein 2003 Answer Key Free and Paid eBooks
 - Evan P Silberstein 2003 Answer Key Public Domain eBooks
 - Evan P Silberstein 2003 Answer Key eBook Subscription Services
 - Evan P Silberstein 2003 Answer Key Budget-Friendly Options
- 6. Navigating Evan P Silberstein 2003 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Evan P Silberstein 2003 Answer Key Compatibility with Devices
 - Evan P Silberstein 2003 Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Evan P Silberstein 2003 Answer Key
 - Highlighting and Note-Taking Evan P Silberstein 2003 Answer Key
 - Interactive Elements Evan P Silberstein 2003 Answer Key
- 8. Staying Engaged with Evan P Silberstein 2003 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Evan P Silberstein 2003 Answer Key
- 9. Balancing eBooks and Physical Books Evan P Silberstein 2003 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Evan P Silberstein 2003 Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Evan P Silberstein 2003 Answer Key
 - Setting Reading Goals Evan P Silberstein 2003 Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Evan P Silberstein 2003 Answer Key
 - Fact-Checking eBook Content of Evan P Silberstein 2003 Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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