

Section One: Multiple-choice

(50 marks)

1. D	2. C	3. B	4. C	5. D
6. C	7. C	8. D	9. B	10. C
11. D	12. B	13. B	14. C	15. A
16. All	17. B	18. D	19. A	20. C
21. C	22. B	23. C	24. D	25. C

Section Two: Short answer

(70 marks)

Question 26

(4 marks)

- (a) Some hydrochloric acid solution is mixed with solid sodium carbonate. (2 marks)

A white solid dissolves in a colourless solution, producing a colourless and odourless gas. (2)

(*Must have two observations for both marks).

- (b) Some solid copper (II) hydroxide is mixed with a dilute nitric acid solution. (2 marks)

A blue solid dissolves in a colourless solution to produce a blue solution. (2)

(*Must have two observations for both marks).

Question 27

(6 Marks)

- (a) Write balanced chemical equations showing carbon dioxide reacting with water to form carbonic acid, and then the two successive ionisation reactions that carbonic acid undergo in water. (3 marks)



- (b) Identify a conjugate acid-base pair in this reaction, and explain why it is classified as a Bronsted – Lowry acid-base reaction. (3 marks)

Conjugate A/B pair = $\text{CO}_3^{2-} / \text{HCO}_3^-$ (1) *Also accept $\text{HCO}_3^- / \text{CO}_3^{2-}$

This equation is classified as a Bronsted – Lowry acid-base reaction because in the forward reaction, H_2O donates a proton, thus acting as a B-L acid, (1)
while CO_3^{2-} accepts a proton, thus acting as a B-L base. (1)

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Linda D. Williams



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Table of Contents Chemistry Semester 1 Exam Answer Key

1. Understanding the eBook Chemistry Semester 1 Exam Answer Key
 - The Rise of Digital Reading Chemistry Semester 1 Exam Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Semester 1 Exam Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Semester 1 Exam Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Semester 1 Exam Answer Key
 - Personalized Recommendations

- Chemistry Semester 1 Exam Answer Key User Reviews and Ratings
- Chemistry Semester 1 Exam Answer Key and Bestseller Lists
- 5. Accessing Chemistry Semester 1 Exam Answer Key Free and Paid eBooks
 - Chemistry Semester 1 Exam Answer Key Public Domain eBooks
 - Chemistry Semester 1 Exam Answer Key eBook Subscription Services
 - Chemistry Semester 1 Exam Answer Key Budget-Friendly Options
- 6. Navigating Chemistry Semester 1 Exam Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Semester 1 Exam Answer Key Compatibility with Devices
 - Chemistry Semester 1 Exam Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Semester 1 Exam Answer Key
 - Highlighting and Note-Taking Chemistry Semester 1 Exam Answer Key
 - Interactive Elements Chemistry Semester 1 Exam Answer Key
- 8. Staying Engaged with Chemistry Semester 1 Exam Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Semester 1 Exam Answer Key
- 9. Balancing eBooks and Physical Books Chemistry Semester 1 Exam Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Semester 1 Exam Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Semester 1 Exam Answer Key
 - Setting Reading Goals Chemistry Semester 1 Exam Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Semester 1 Exam Answer Key
 - Fact-Checking eBook Content of Chemistry Semester 1 Exam Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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