

CHEMISTRY

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PERIODIC TABLE OF THE ELEMENTS

The periodic table is color-coded by groups:

- 1 (pink):** s-subshell, 1 orbital
- 2 (orange):** s-subshell, 1 orbital
- 3 (light blue):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 4 (medium blue):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 5 (dark blue):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 6 (teal):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 7 (light green):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 8 (medium green):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 9 (dark green):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 10 (light purple):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 11 (medium purple):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 12 (dark purple):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 13 (pinkish-purple):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 14 (pink):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 15 (orange-pink):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 16 (orange):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 17 (light orange):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 18 (yellow):** s-subshell, 1 orbital; p-subshell, 3 orbitals

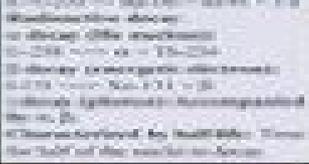


ATOMIC STRUCTURE

Atomic Number, Z , is of protons in the nucleus. For neutral atom, $Z = +$ of electrons. For ions of charge, $+Z$ (cations) or $-Z$ (anions). Mass Number, A , is $Z + N$, the number of nucleons in the nucleus. Isotopes atoms with the same Z , different N , the changed sign have a number of protons, say Z , and a different number of neutrons, as indicated with the symbol. Mass Number or Atomic Weight is printed above of element boxes for a given element.

NUCLEAR CHEMISTRY

Nuclear reactions with the nucleus particles consist of chemical behavior change and mass energy change. $E = mc^2$ relates to mass loss, known as the Einstein equation. $E = mc^2$.
Nucleus is built from protons, Z , and is symbolized by ${}^A_Z X$, or ${}^A_Z \text{Element}$, Z or AZ .
Types of processes:
Transmutation: Two elements made from particle collisions.
Radioactive decay: α -decay: ${}^4_2\text{He}$, β -decay: ${}^0_{-1}\text{e}$, γ -decay: ${}^0_0\gamma$.
Radioactive decay:
- α -decay: ${}^A_Z X \rightarrow {}^{A-4}_{Z-2} Y + {}^4_2\text{He}$
- β -decay: ${}^A_Z X \rightarrow {}^A_{Z+1} Y + {}^0_{-1}\text{e}$
- γ -decay: ${}^A_Z X^* \rightarrow {}^A_Z X + \gamma$
Characterized by half-life, time for half of the nucleus to decay.



ATOMIC QUANTUM NUMBERS AND ORBITALS

Quantum numbers describe the energy, size, and shape of an orbital. The principal quantum number, n , is the most important. It is the number of shells. $n = 1, 2, 3, \dots$. The angular momentum quantum number, l , is the number of subshells. $l = 0, 1, 2, \dots, n-1$. The magnetic quantum number, m_l , is the number of orbitals in a subshell. $m_l = -l, -(l-1), \dots, 0, \dots, (l-1), l$. The spin quantum number, m_s , is the number of electrons in an orbital. $m_s = +\frac{1}{2}, -\frac{1}{2}$.
Orbitals: $n=1, l=0, m_l=0$ (1 orbital); $n=2, l=0, m_l=0$ (1 orbital); $n=2, l=1, m_l=0$ (1 orbital); $n=2, l=1, m_l=1$ (2 orbitals); $n=3, l=0, m_l=0$ (1 orbital); $n=3, l=1, m_l=0$ (1 orbital); $n=3, l=1, m_l=1$ (2 orbitals); $n=3, l=2, m_l=0$ (1 orbital); $n=3, l=2, m_l=1$ (2 orbitals); $n=3, l=2, m_l=2$ (2 orbitals).

MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe all many-electron atoms and ions. A many electron atom has more electrons than energy levels and orbital properties. The main point is that orbitals have different energies. The filling of levels is guided by the Aufbau Principle.

n	l	m _l	orbitals
1	0	0	1 orbital
2	0	0	1 orbital
2	1	0	1 orbital
2	1	1	2 orbitals
3	0	0	1 orbital
3	1	0	1 orbital
3	1	1	2 orbitals
3	2	0	1 orbital
3	2	1	2 orbitals
3	2	2	2 orbitals

Atomic configuration: Aufbau orbital sequence. Pauli Exclusion Principle: Each orbital has a unique set of quantum numbers. An orbital may hold up to two electrons, one with spin up, $+\frac{1}{2}$, and one with spin down, $-\frac{1}{2}$. Hund's Rule: Electrons fill orbitals singly, as a maximum number of unpaired electrons. Aufbau Principle: Orbitals are filled in order of increasing energy. Aufbau Principle: Orbitals are filled in order of increasing energy. Aufbau Principle: Orbitals are filled in order of increasing energy.

Energy level diagram showing the relative energies of atomic orbitals for different principal quantum numbers (n=1 to n=7). The diagram illustrates how the energy of orbitals increases with n and how subshells (s, p, d, f) split in energy within each shell.

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