

CHEMISTRY MULTIPLE CHOICE QUESTIONS

A. Atoms, molecules and stoichiometry

2002 -2019

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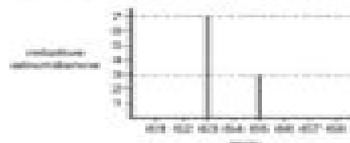
As a consequence, an adult human can be considered to have a daily diet of 1.80 kg of carbohydrate (empirical formula $C_6H_{10}O_5$).

Which mass of carbon dioxide does a person produce each day if all the carbohydrate eaten is digested and oxidised?

- A 0.907 kg B 0.908 kg C 1.70 kg D 3.54 kg

(2002 NA/12)

The diagram shows the mass spectrum of a sample of naturally-occurring copper.



What is the relative atomic mass of this copper?

- A 63.2 B 63.5 C 63.8 D 64.0

(2002 NA/12)

Which pairs of compounds have the same empirical formula?

- ethane and ethene
- ethane and cyclohexane
- cyclohexane and oct-1-ene

(2002 NA/13)

A mixture of 1.00 dm^3 of methane and 1.00 dm^3 of ethane was sparked with an excess of oxygen. After cooling to room temperature, the residual gas was passed through aqueous potassium hydroxide.

Which compound has the empirical formula CH_2O ?

- methanol
- ethanoic acid
- methyl methanoate

(2002 NA/14)

What is the number of molecules in 500 cm^3 of oxygen under room conditions?

- A 1.25×10^{23}
 B 1.25×10^{24}
 C 3.0×10^{23}
 D 3.0×10^{24}

(2004 MA/12)

In the preparation of soft margarine, glycerol trioleate



is fully hydrogenated so that, on average, one of its side-chains is converted into the $\text{CH}_3(\text{CH}_2)_7\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$ moiety and two side-chains are converted into the $\text{CH}_3(\text{CH}_2)_7\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$ moiety.

How many moles of hydrogen are required to convert one mole of glycerol trioleate into the soft margarine?

- A 4 B 5 C 6 D 8

(2008 MA/12)

Use of the Data Booklet is relevant to this question.

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JA Banks



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Table of Contents Chemistry Multiple Choice Questions And Answers

1. Understanding the eBook Chemistry Multiple Choice Questions And Answers
 - The Rise of Digital Reading Chemistry Multiple Choice Questions And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Multiple Choice Questions And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Multiple Choice Questions And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Multiple Choice Questions And Answers
 - Personalized Recommendations
 - Chemistry Multiple Choice Questions And Answers User Reviews and Ratings
 - Chemistry Multiple Choice Questions And Answers and Bestseller Lists

5. Accessing Chemistry Multiple Choice Questions And Answers Free and Paid eBooks
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 - Highlighting and Note-Taking Chemistry Multiple Choice Questions And Answers
 - Interactive Elements Chemistry Multiple Choice Questions And Answers
8. Staying Engaged with Chemistry Multiple Choice Questions And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Multiple Choice Questions And Answers
9. Balancing eBooks and Physical Books Chemistry Multiple Choice Questions And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Multiple Choice Questions And Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Multiple Choice Questions And Answers
 - Setting Reading Goals Chemistry Multiple Choice Questions And Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Multiple Choice Questions And Answers
 - Fact-Checking eBook Content of Chemistry Multiple Choice Questions And Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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