

Chemistry Notes 14.3

Three Basic Assumptions of the kinetic theory:

- 1) Gas is composed of particles, usually molecules or atoms
- 2) Particles in a gas move rapidly in constant random motion (they move in straight line paths only colliding with one another or other objects)
- 3) Collisions are perfectly elastic: kinetic energy is transferred without loss from one particle to another

Equation: $P \times V = n \times R \times T$

$R = 8.31 \text{ (L} \times \text{kPa) / (mol} \times \text{K)}$

$n = (P \times V) / (R \times T)$

Ideal Gas:

- 1) A gas is composed of particles with a mass that is low enough to be considered almost zero.
 - But a real gas has limitations to the pressure it can be compressed to before it condenses into a liquid

Real gases differ from ideal gases in the following ways:

- Real gases have small attractive and repulsive forces between particles, and ideal gases do not
- Real gas particles have a volume, and ideal gas particles do not
- Real gas particles collide inelastically and lose energy with collisions, and ideal gas particles collide elastically
- Real gases are not infinitely compressible

Ideal gases don't exist because they do take up space, and there are attractive forces

Diffusion: molecules moving from areas of high concentration to low concentration

- Perfume

Effusion: Gas escaping through tiny hole in container

- Both of these depend on the molar mass of the particle, which determines the speed
- The rate of effusion and diffusion is inversely proportional to the square root of the molar mass

To solve Graham's law problems, put whatever you're looking for in the numerator

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Table of Contents Chemistry Ch 14 Guided Reading

1. Understanding the eBook Chemistry Ch 14 Guided Reading
 - The Rise of Digital Reading Chemistry Ch 14 Guided Reading
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Ch 14 Guided Reading
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Ch 14 Guided Reading
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Ch 14 Guided Reading
 - Personalized Recommendations
 - Chemistry Ch 14 Guided Reading User Reviews and Ratings
 - Chemistry Ch 14 Guided Reading and Bestseller Lists
5. Accessing Chemistry Ch 14 Guided Reading Free and Paid eBooks
 - Chemistry Ch 14 Guided Reading Public Domain eBooks
 - Chemistry Ch 14 Guided Reading eBook Subscription Services

- Chemistry Ch 14 Guided Reading Budget-Friendly Options
- 6. Navigating Chemistry Ch 14 Guided Reading eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Ch 14 Guided Reading Compatibility with Devices
 - Chemistry Ch 14 Guided Reading Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Ch 14 Guided Reading
 - Highlighting and Note-Taking Chemistry Ch 14 Guided Reading
 - Interactive Elements Chemistry Ch 14 Guided Reading
- 8. Staying Engaged with Chemistry Ch 14 Guided Reading
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Ch 14 Guided Reading
- 9. Balancing eBooks and Physical Books Chemistry Ch 14 Guided Reading
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Ch 14 Guided Reading
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Ch 14 Guided Reading
 - Setting Reading Goals Chemistry Ch 14 Guided Reading
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Ch 14 Guided Reading
 - Fact-Checking eBook Content of Chemistry Ch 14 Guided Reading
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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