

ChemActivity 40 B

Relationships among Equilibrium Constants

We saw in ChemActivity 40 that equilibrium constants can be calculated if similar reactions have known values of K .

1. Given that $2\text{NO}(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$; $K_p = 1.7 \times 10^{17}$ at 25°C

(a) Write the equilibrium expression $K_p =$

A similar reaction for which we can calculate K is: $\text{NO}(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightleftharpoons \text{NO}_2(\text{g})$.

(b) Write the equilibrium expression for the second reaction, $K_p' =$

(c) Describe the mathematical relationship between K_p and K_p' (how is K_p related to K_p' ?)

(d) What is the numerical value of K_p' ?

A third similar reaction is $\text{NO}_2(\text{g}) \rightleftharpoons \text{NO}(\text{g}) + \frac{1}{2}\text{O}_2(\text{g})$

(e) Write the equilibrium expression for the third reaction, $K_p'' =$

(f) Describe the mathematical relationship between K_p' and K_p'' .

(g) What is the numerical value of K_p'' ?

In this course we have had need to add chemical reaction in a mathematical way. In our study of thermochemistry we added chemical reactions in applying Hess's Law. In each case all the reactants and products are combined and compounds that appear as both reactants and products are "canceled" from the final (overall) reaction. The same can be done with equilibrium expressions.

If we want to know the equilibrium constant for the atmospheric reaction



We can use reactions of known K that sum to give the desired equilibrium.

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