

Pre-lab Questions for Experiment #5

Rates of Chemical Reactions I: A Clock Reaction

1) Complete and balance the following reactions:



2) The rate law for the reaction



is $\text{rate} = k[H_2][NO]^2$. Which of the following mechanisms can be ruled out on the basis of the observed rate expression?

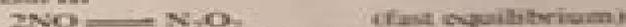
Mechanism I



Mechanism II



Mechanism III



3) Consider the reaction



From the following data obtained at a certain temperature, determine the order of the reaction and calculate the rate constant:

[A] (M)	[B] (M)	Rate (M/s)
1.50	1.50	3.20×10^{-1}
1.50	2.50	3.20×10^{-1}
3.00	1.50	6.40×10^{-1}

4) The following gas-phase reaction was studied at 290°C by observing the change in pressure as a function of time in a constant-volume vessel:



Determine the order of the reaction and the rate constant based on the following data:

Time (s)	Partial Pressure of $ClCO_2CCl_3$ (mmHg)
0	15.76
181	12.64
513	8.73
1164	4.44

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The Chemical News and Journal of Industrial Science, 1919 *American Medical Directory*, 1912 **Question and Answer Review for Phlebotomy** Kathleen Becan-McBride, Diana Garza, 2000-11 Prentice Hall Health Review Series

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About the Authors Kathleen Becan McBride EdD MT ASCP CLS NCA is the Director of Community Outreach Education at the University of Texas Health Science Center at Houston She has published 13 books and over 50 journal articles She has been educating students in the health professions since 1973 Diana Garza EdD MT ASCP CLS NCA is Associate Research Professor and Project Director for Interdisciplinary Health Care Education at Texas Woman s University Institute of Health Sciences in Houston Texas She is also an Adjunct Clinical Associate Professor in the Department of Laboratory Medicine at the University of Texas M D Anderson Cancer Center in Houston She has published nine books numerous journal articles and has been educating students in the health professions since 1981

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