

CH-101 2008 Assignment 10 Class

Name: _____

Number: _____

10. Atmospheric pressure at elevations of 1000 ft average about 0.70 atm. Consider this pressure to maintain air masses, lakes of mercury, and gas-filled and vacuum sealed flasks at equilibrium at 1000 feet mean sea level (Mount St. Helens)?

What is the physical explanation regarding water pressure in commercial scuba tanks provided constructed to a pressure greater than the required hydrostatic pressure at construction of 1000 ft?

11. Suppose a 1000 mL volume of water with a volume of 2.0 L, from surface, during the process at 1.0 atm, has a depth of 200 m, where the pressure is 3.0 atm. What would happen to the volume of the balloon? What if the end of the balloon was in a temperature that went to the surface and was returned to surface? What would happen to the air the other direction?
12. A volume of gas is provided to a cylinder with a freely movable piston, and then it is pressurized then decompressed to return to that initial configuration. Consider the amount of gas, volume of the cylinder, temperature to mean daily for year?

- (a) The pressure will reduce to half that the volume will double.
- (b) The pressure will reduce to half the volume will double.
- (c) The pressure will reduce to half that the volume will reduce that number.
- (d) The pressure will increase that something that pressure will double.

13. Use Charles's law to compare the volume pressure and amount of compressed gas for the conditions:

1	2	3	4
1 atm	100°C	1 L	1000 mol
1	100°C	1000 mol	1 atm
1000 mol	100°C	1000 mol	1000 mol
1000 mol	100°C	1000 mol	1000 mol

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Proceedings of the 1995 ERDEC Scientific Conference on Chemical and Biological Defense Research, 14-17

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