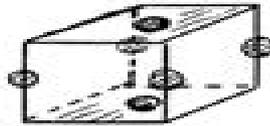


PROBLEM SOLUTIONS
CHAPTER 1

(1.1)
(a)



... $\frac{1}{2}$ atom inside cell on top and bottom faces
... $\frac{1}{4}$ atom along each vertical edge

$$\text{TOTAL ATOMS} = 2 \times \frac{1}{2} + 4 \times \frac{1}{4} = \underline{\underline{2}}$$

(b)



⇐ TOP VIEW

- ... plane #1
- △ ... plane #2
- ... plane #3

(c) bcc

(1.2)

(a) The diamond lattice is composed of two interpenetrating fcc lattices. Each fcc lattice unit cell contains 4 atoms. ($\frac{1}{2}$ atom at each of 6 faces and $\frac{1}{8}$ atom at each of 8 corners.) Thus the Si lattice contains 8 ATOMS per unit cell. Inspecting Fig. 1.5a leads to the same conclusion -- there are 4 atoms totally inside the cube and an fcc arrangement on the faces of the cube.

(b)

$$\frac{\text{Si ATOMS}}{\text{cm}^3} = \frac{\text{ATOMS/UNIT CELL}}{a^3} = \frac{8}{(5.431 \times 10^{-8})^3} = 4.997 \times 10^{22} \checkmark$$

↙ from Table 1.5

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SJ Ball



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