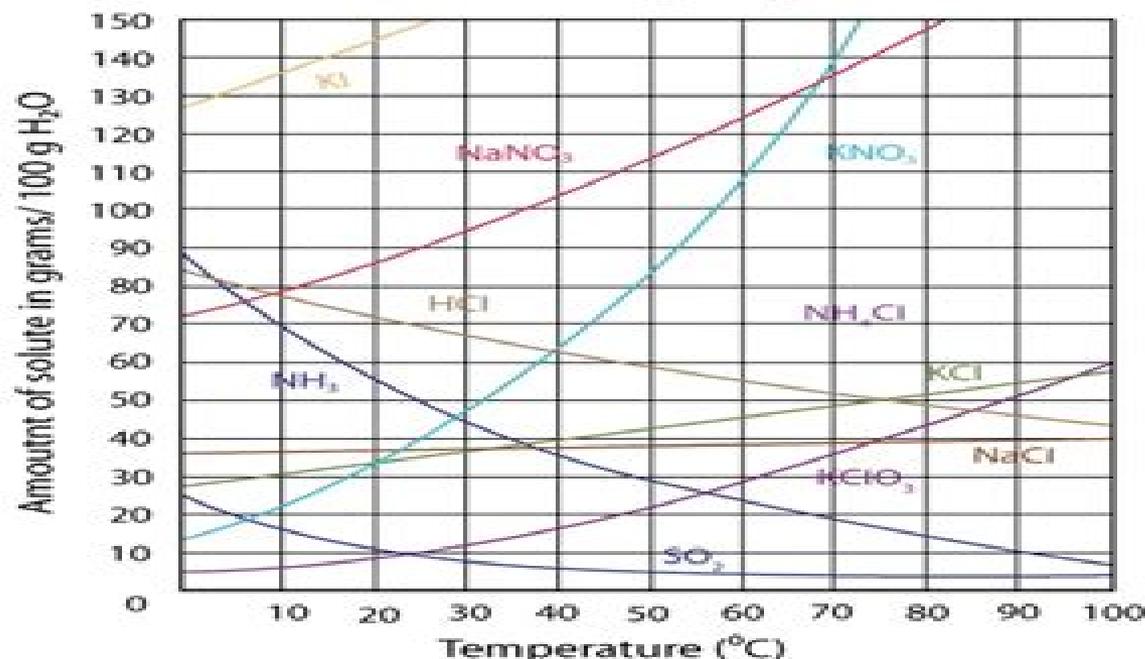




# Interpreting Solubility Curve



1. Define solubility. \_\_\_\_\_
2. Using the solubility curve below, answer the given questions.



- I. What is the solvent in this solubility curve? \_\_\_\_\_
- II. Saturation points are given per how much of the solvent? \_\_\_\_\_
- III. What is the saturation point of NaNO<sub>3</sub> at 10°C, 40°C, & 80°C?  
 10°C = \_\_\_\_\_      40°C = \_\_\_\_\_      80°C = \_\_\_\_\_
- IV. How much NH<sub>3</sub> can you dissolve in water at 10°C, 30°C, & 90°C?  
 10°C = \_\_\_\_\_      30°C = \_\_\_\_\_      90°C = \_\_\_\_\_
- V. Which salt is the least soluble in water? \_\_\_\_\_
- VI. How many grams of potassium chloride can be dissolved in 200 g of water at 80 °C? \_\_\_\_\_
- VII. At 40 °C, how much potassium nitrate will be dissolved in 300 g of water? \_\_\_\_\_
- VIII. Which salt shows the least change in solubility from 0° to 100° C? \_\_\_\_\_
- IX. Which substance shows a decrease in solubility from 0° to 100° C? \_\_\_\_\_
- X. At 30 °C, 90 g of sodium nitrate is dissolved in 100 g water. Is this solution saturated, unsaturated, or supersaturated? \_\_\_\_\_

# Solubility Temperature Graphs Chapter 14 Worksheet Answers

**Rachel Sandford**



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