

Student Exploration: Solubility and Temperature

Vocabulary:

1. concentration - Molar concentration is a measure of the concentration of a chemical species, in particular of a solute in a solution, in terms of amount of substance per unit volume of solution.
2. dissolve - dissolve is to cause a solute to pass into a solution.
3. homogeneous mixture - A homogeneous mixture is a solid, liquid, or gaseous mixture that has the same proportions of its components throughout any given sample.
4. solubility - Solubility is the property of a solid, liquid or gaseous chemical substance called solute to dissolve in a solid, liquid or gaseous solvent.
5. solubility curve - A solubility curve is a data based graph comparing the amount of solute that will dissolve in a given amount of solvent at various temperatures.
6. solute - the minor component in a solution, dissolved in the solvent.
7. solution - a liquid mixture in which the minor component (the solute) is uniformly distributed within the major component (the solvent).
8. solvent - able to dissolve other substances.

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. What happens when you stir a spoonful of sugar into hot water?
The sugar dissolves into the water.
2. When sugar or another substance is **dissolved** in water, it disappears from view and forms a **homogeneous mixture** with the water, also called a **solution**. The water tastes sweet and may appear slightly cloudy.
If you can't see the sugar, how can you tell that it is there?
3. Does sugar dissolve more easily in hot water or cold water? **Hot water**

Gizmo Warm-up

A solution generally consists of two parts, a **solute** that is dissolved and a **solvent** that the solute is dissolved into. For example, sugar is a solute that is dissolved into the solvent water. In the *Solubility and Temperature* Gizmo, you will study how temperature affects how much solute will dissolve in a solution.

To begin, check that **Potassium nitrate** is selected and the **Temp.** of the water is 20 °C. Click **OK**.

1. In this solution, what is the solute? **Potassium nitrate** What is the solvent? **Water**
2. Click **Add 10 g** to mix 10 g of potassium nitrate into the water.
 - A. Did all of the potassium nitrate dissolve? **yes**
 - B. How can you tell? **All of the potassium nitrate disappeared as it was stirred into the water.**



Solubility And Temperature Answers Key

Peter Ellis, Derek McMonagle



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