

KINETICS Practice Problems and Solutions

Name: _____
 Period: _____ Date: _____

AP Chemistry
 Dr. Mandes

The following questions represent potential types of quiz questions. Please answer each question completely and thoroughly. The solutions will be posted on-line on Monday.

5. Please do #18 in chapter 12 of your text.
- This increases the $[H_2]$, which will increase the rate, but has no effect on k
 - Due to Arrhenius equation, changing temperature changes the value of k .
 - Catalyst function by lowering the activation energy, so due to Arrhenius equation, changing the activation energy changes the value of k .

6. Consider the reaction: $P_4 + 6 H_2 \rightarrow 4 PH_3$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[P_4]$, mol/L	$[H_2]$, mol/L	Initial Rate, mol/L · s
0.0110	0.0075	3.20×10^{-4}
0.0110	0.0150	6.40×10^{-4}
0.0220	0.0150	6.39×10^{-4}

- What is the order with respect to: P_4 0
 H_2 1
- Write the rate law for this reaction. $rate = k[H_2]$
- Determine the value and units of the rate constant, k . **plug and chug using the rate law & data from exp't 1 and solving for k , we get $k = 0.0427 s^{-1}$**

7. Consider the reaction: $SO_2 + O_3 \rightarrow SO_3 + O_2$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[SO_2]$, mol/L	$[O_3]$, mol/L	Initial Rate, mol/L · s
0.25	0.40	0.118
0.25	0.20	0.118
0.75	0.20	1.062

- What is the order with respect to: SO_2 2
 O_3 0
- Write the rate law for this reaction. $rate = k[SO_2]^2[O_3]^0$
- Determine the value and units of the rate constant, k . **plug and chug using the rate law & data from exp't 1 and solving for k , we get $k = 2.36 molL^{-2} s^{-1}$**

8. Consider the following mechanism.
- $$A_2 + B_2 \rightarrow R + C \quad (\text{slow})$$
- $$A_2 + R \rightarrow C \quad (\text{fast})$$

- Write the overall balanced chemical equation. $2 A_2 + B_2 \rightarrow 2 C$
- Identify any intermediates within the mechanism. **R**
- What is the order with respect to each reactant? **A_2 : 1st; B_2 : 1st**

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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields. *Technical Mechanics, Statics, Kinematics, Kinetics* Edward Rose Maurer, Raymond Jefferson Roark, 1925 [Super Course for the MCAT](#) Stefan Bosworth, 1992-03 This brand new addition to Arco's acclaimed SuperCourse series is certain to set a new high standard for quality. It offers the comprehensive MCAT preparation including in depth subject reviews as well as full length sample exams and in depth preparation for the new required essay section.

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College of Engineering Catalogue, Carnegie Institute of Technology, College of Engineering and Science, 1927

Kinetics and Catalysis, 1983-07 **Cracking the GRE Chemistry Test**, 2002. The Princeton Review realizes that acing the GRE Chemistry Test is very different from getting straight A's in school. We don't try to teach you everything there is to know about chemistry; only the techniques you'll need to score higher on the test. There's a big difference. In **Cracking the GRE Chemistry Test**, we'll teach you how to think like the test writers and eliminate answer choices that look right but are planted to fool you. Raise your score by focusing on the material most likely to appear on the test. Test your knowledge with review questions for each chemistry topic covered. This book includes one full-length practice GRE Chemistry Test. All of our practice questions are like the ones you'll see on the actual GRE Chemistry Test, and we fully explain every solution.

The Ohio Journal of Science, 1959. Includes book reviews and abstracts. **The Software Directory for the APPLE Computer**, 1981. **Biochemical Calculations**, Irwin H. Segel, 1968. Weak acids and bases, Amino acids and peptides, Biochemical energetics, Enzyme kinetics, Spectrophotometry, Isotopes in biochemistry, Miscellaneous calculations.

Student's Guide to Chemistry; a Modern Introduction, David Brooks, 1974. Fundamentals of Engineering, Merle C. Potter, 1999. Muskingum College Bulletin, Muskingum College, 1924. **A Problem-Solving Approach to Aquatic Chemistry**, James N. Jensen, 2003-03-13. This text provides a detailed introduction to aquatic equilibrium chemistry, calculation methods for systems at equilibrium, applications of aquatic chemistry, and chemical kinetics. Software designed especially for the text allows the reader to build complex models by applying equilibrium calculation principles. Important features include material-specific and integrated case studies, thought-provoking questions, key ideas, and historical sketches.

Elementary Mechanics, United States Naval Institute, 1922. **A Text-book of Physical Chemistry, Theory and Practice**, Arthur Woolsey Ewell, 1909. **The Gas Record**, 1915.

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