

3 • Reactions in Aqueous Solution**STUDY QUESTIONS AND PROBLEMS**

1. Classify each of the following solutes as a strong electrolyte, weak electrolyte, or nonelectrolyte:

sugar	sodium hydroxide
common salt (NaCl)	hydrochloric acid
alcohol	copper sulfate
acetic acid	carbonic acid

2. Predict the solubility of the following salts:

- sodium sulfate
- potassium chromate
- silver bromide
- nickel (II) hydroxide
- aluminum nitrate
- barium sulfide
- ammonium acetate
- strontium iodide

3. Write the ions that are produced when the following substances dissolve in water:

- $\text{Mg}(\text{OH})_2$
- K_2SO_4
- NaHCO_3
- $(\text{NH}_4)_3\text{PO}_4$
- NaClO

4. Predict whether or not the following reactions will lead to a precipitate. Write complete and net ionic equations for all the reactions.

- a. potassium chromate and lead acetate
- b. silver perchlorate and ammonium chloride
- c. potassium carbonate and copper acetate
- d. sodium fluoride and magnesium iodide
- e. barium nitrate and potassium sulfate

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