

### Basic Stoichiometry Post-Lab Homework Exercises

1. Load the "Reactants, Products, and Leftovers" simulation and work through each of the levels of the Game! At <http://phet.colorado.edu/> or googling "phet." You may have to download or update Java on your computer.

**Show your work for all calculations. Remember: Labels, units, sig figs and cancellation!**

2. For the reaction  $P_4 + 6Cl_2 \rightarrow 4PCl_3$ , determine how many moles of chlorine  $Cl_2$  would be needed to react with 4.0 moles of phosphorus  $P_4$  to entirely use up all the phosphorus. 4) \_\_\_\_\_

3. If 3.00 moles of  $P_4$  reacted with 22.0 moles  $Cl_2$  according to the above reaction, determine:

- a) How many moles  $PCl_3$  are produced a) \_\_\_\_\_
- b) How many moles of  $P_4$  are left in excess after the reaction (if any) b) \_\_\_\_\_
- c) How many moles of  $Cl_2$  are left in excess after the reaction (if any) c) \_\_\_\_\_

In reality, reactants don't have to react in perfect whole-numbers of moles. In a two-reactant synthesis reaction, usually **one reactant gets entirely used up (and determines how much product is made)**, even if that means using fractions of a mole of reactant. For instance, when solid, metallic aluminum  $Al$  and red, liquid bromine  $Br_2$  are brought together, they make a white solid according to the reaction  $2Al + 3Br_2 \rightarrow 2AlBr_3$ . If 5.0 moles of aluminum  $Al$  was reacted with 10 moles bromine  $Br_2$ , all five moles of aluminum would react, with only 7.5 moles bromine. (2:3 mole ratio) This would produce only 5.0 moles of  $AlBr_3$ , leaving 2.5 moles of excess  $Br_2$  behind.

5. Now assume 3.0 moles  $Al$  and 4.0 moles  $Br_2$  react.

- a) How much (in moles)  $AlBr_3$  gets produced? a) \_\_\_\_\_
- b) Which chemical must be the *excess reactant*? b) \_\_\_\_\_
- c) Which chemical is the *limiting reactant*? c) \_\_\_\_\_
- d) If all the limiting reactant gets used up, how much of the excess reactant is left? d) \_\_\_\_\_

# Basic Stoichiometry Post Lab Homework Exercises Answers

**AW Chickering**



**Basic Stoichiometry Post Lab Homework Exercises Answers:**

## Unveiling the Magic of Words: A Review of "**Basic Stoichiometry Post Lab Homework Exercises Answers**"

In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their ability to kindle emotions, provoke contemplation, and ignite transformative change is really awe-inspiring. Enter the realm of "**Basic Stoichiometry Post Lab Homework Exercises Answers**," a mesmerizing literary masterpiece penned by way of a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve to the book is central themes, examine its distinctive writing style, and assess its profound impact on the souls of its readers.

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