

## Dimensional Analysis Practice Worksheet Solutions

Complete this and place it in your portfolio after section 2-2. I will email the solutions later so that you can check your work (make sure you are showing the work the same way you see it in the solutions, using dimensional analysis). It is important that you practice and learn to use dimensional analysis (multiplying by unit fractions) appropriately. Make sure you get help if you are still having difficulties with this topic. Use my office hours or the math tutors in the Student Learning Center in the library. You do not have to write sentences on this practice worksheet. You do need to show all of the units and use  $\approx$  for approximations.

Do the following conversions using dimensional analysis using the conversion factors on page 74 of your text. You should be doing these all in one computation (even if you have to use more than one unit fraction), not several computations. I have done #1 as example.

1. 5400 inches to miles

$$5400 \cancel{\text{in.}} \times \frac{1 \cancel{\text{ft}}}{12 \cancel{\text{in.}}} \times \frac{1 \text{ mi}}{5280 \cancel{\text{ft}}} = \frac{5400}{63360} \text{ mi} \\ \approx 0.085 \text{ mi}$$

2. 13 weeks to seconds

$$13 \cancel{\text{ weeks}} \times \frac{7 \cancel{\text{ days}}}{1 \cancel{\text{ week}}} \times \frac{24 \cancel{\text{ hrs}}}{1 \cancel{\text{ day}}} \times \frac{60 \cancel{\text{ min}}}{1 \cancel{\text{ hr}}} \times \frac{60 \text{ sec}}{1 \cancel{\text{ min}}} = 7,862,400 \text{ sec}$$

3. 54 yards to mm

$$54 \cancel{\text{ yards}} \times \frac{3 \cancel{\text{ feet}}}{1 \cancel{\text{ yd}}} \times \frac{1 \cancel{\text{ m}}}{3.28 \cancel{\text{ ft}}} \times \frac{1000 \text{ mm}}{1 \cancel{\text{ m}}} = \frac{162000}{3.28} \text{ mm} \\ \approx 49,390.24 \text{ mm}$$

4.  $36 \frac{\text{cm}}{\text{sec}}$  to mph

$$\frac{36 \cancel{\text{ cm}}}{1 \cancel{\text{ sec}}} \times \frac{60 \cancel{\text{ sec}}}{1 \text{ hr}} \times \frac{1 \cancel{\text{ m}}}{1000 \cancel{\text{ cm}}} \times \frac{3.28 \cancel{\text{ ft}}}{1 \cancel{\text{ m}}} \times \frac{1 \text{ mi}}{5280 \cancel{\text{ ft}}} = \frac{7084.8 \text{ mi}}{5280000 \text{ hr}} \\ \approx 0.00134 \frac{\text{mi}}{\text{hr}}$$

# Chemistry Dimensional Analysis Practice Problems Answers

**James P. Birk, Donald M. Baird**



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**Ebook: Chemistry** Julia Burdge, 2014-10-16 Chemistry Third Edition by Julia Burdge offers a clear writing style written with the students in mind Julia uses her background of teaching hundreds of general chemistry students per year and creates content to offer more detailed explanation on areas where she knows they have problems With outstanding art a consistent problem solving approach interesting applications woven throughout the chapters and a wide range of end of chapter problems this is a great third edition text

**Chemistry** Jason Overby, Raymond Chang, 2024 The fifteenth edition continues a long tradition of providing a firm foundation in the concepts of chemical principles while instilling an appreciation of the important role chemistry plays in our daily lives We believe that it is our responsibility to assist both instructors and students in their pursuit of this goal by presenting a broad range of chemical topics in a logical format At all times we strive to balance theory and application and to illustrate principles with applicable examples whenever possible

**NAPLEX 2016 Strategies, Practice, and Review with 2 Practice Tests** Amie D. Brooks, Cynthia Sanoski, Emily R. Hajjar, Brian R. Overholser, 2016-03-29 Kaplan s NAPLEX 2016 Strategies Practice and Review with 2 Practice Tests is a step by step guide to scoring higher on the North American Pharmacist Licensure Examination Features NEW Chapters on renal disorders hepatic disorders transplant immunosuppression and preventive medicine NEW 50 percent more end of chapter review questions to assess your content knowledge 200 comprehensive tables summarizing must know drugs at a glance including the 200 most prescribed drugs 2 full length practice tests Online practice test that simulates the exam experience and actively tracks your areas of strength and weakness Robust calculations section with 30 practice problems and step by step math solutions Organized streamlined coverage of each NAPLEX subject area Suggested study times for each NAPLEX

topic to help calibrate your study plan

**Chemistry: Matter & Change, Solving Problems - A Chemistry Handbook** McGraw Hill, 2001-08  
**Glencoe Chemistry Solving Problems A Chemistry Handbook Matter and Change** Prentice Hall  
**Chemistry**, 2000 *Foundations of College Chemistry* Morris Hein, 1991 Features strong problem solving and carefully constructed problem sets

**Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl** Paul B. Kelter, 1993

**Introduction to General, Organic, and Biochemistry** Morris Hein, Scott Pattison, Susan Arena, Leo R. Best, 2014-01-15  
The most comprehensive book available on the subject  
**Introduction to General Organic and Biochemistry 11th Edition** continues its tradition of fostering the development of problem solving skills featuring numerous examples and coverage of current applications Skillfully anticipating areas of difficulty and pacing the material accordingly this readable work provides clear and logical explanations of chemical concepts as well as the right mix of general chemistry organic chemistry and biochemistry An emphasis on real world topics lets readers clearly see how the chemistry will apply to their career

**General Chemistry** James P. Birk, Donald M. Baird, 1994-03 Preparatory Chemistry Howard Stephen Stoker, 1990  
Dosage and Solutions Calculations Alice J. Smith, 1989 Chemistry George Tyler Miller, 1987

**General Chemistry, Reactions First** Kevin Revell, 2024-12-04  
Revell's General Chemistry empowers students to grasp essential topics and concepts with more ease Using a friendly approach the text uses metaphors and relatable examples to demystify even the most challenging subjects in general chemistry

Understanding Chemistry Fred M. Dewey, 1994 *Chemistry, Student Solutions Manual* James E. Brady, Joel W. Russell, John R. Holum, 2000-01-26  
The images on the cover call attention to the relationship between macro observations and the intimate structure of chemical substances and the changes both chemical and physical that they undergo

**Fireworks** One of the ingredients is phosphorus a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light

**Carbon** The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball shaped molecules often referred to as buckeyballs Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth

**Acetylene torch** Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen a reaction between molecules of  $C_2H_2$  and  $O_2$  to give  $CO_2$  and  $H_2O$

**Hot air balloon** The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and eight hydrogen atoms

**Stormy weather** The evaporation of water serves to store energy provided by the sun Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet

Chemistry Theodore Lawrence Brown, 1988

**Chemistry** John W. Moore, Conrad L. Stanitski, Peter C. Jurs, 2002  
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