

3. (i) Mole fraction: Commonly used symbol for mole fraction is x and subscript used on the right hand side of x denotes the component. It is defined as:

$$\text{Mole fraction of a component} = \frac{\text{Number of moles of the component}}{\text{Number of moles of all the component}}$$

- (ii) Molality: Molality (m) is defined as the number of moles of the solute per kilogram (kg) of the solvent and is expressed as:

$$\text{Molarity (m)} = \frac{\text{Moles of solute}}{\text{Mass of solvent in kg}}$$

- (iii) Molarity: Molarity (M) is defined as number of moles of solute dissolved in one litre (or one cubic decimetre) of solution.

$$\text{Molarity (M)} = \frac{\text{Moles of solute}}{\text{Mass of solution in litre}}$$

- (iv) Mass percentage: The mass percentage of a component of a solution is defined as:

$$\text{Mass \% of a component} = \frac{\text{Mass of component in the solution}}{\text{Total mass of the solution}} \times 100$$

4. Concentrated nitric acid used in laboratory work is 68% nitric acid by mass in aqueous solution. What should be the molarity of such a sample of the acid if the density of the solution is 1.504 g mL^{-1} ?

4. Given:

Concentration of Nitric Acid, $\text{HNO}_3 = 68\%$

Density of solution, $d = 1.504 \text{ g/ml}$

To find: Molarity, M_o

Formula:

$$\text{Density, } d = \frac{\text{Mass [M]}}{\text{Volume [V]}}$$

$$\text{Molarity, } M_o = \frac{\text{(Number of moles of solute)}}{\text{(Volume of solution in litres)}}$$

68% of Nitric acid by mass in aqueous solution means that 68g $[(68 \times 100)/100]$ of Nitric acid present in 100g of solution.

$$\Rightarrow \text{Molecular mass of Nitric Acid, } \text{HNO}_3 = \frac{[1 \times 1] + [1 \times 14] + [16 \times 3]}{= 63\text{g}}$$

$$\Rightarrow \text{Number of moles of Nitric Acid} = \frac{[68/63]}{= 1.079 \text{ moles}}$$

$$\Rightarrow \text{Given Density, } d = 1.504 \text{ g/ml}$$

$$\Rightarrow \text{Volume, } v = \frac{[100/1.504]}{= 66.489 \text{ ml}}$$

$$\Rightarrow \text{Molarity, } M_o = \frac{[1.079/66.489] \times 1000}$$

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**Daniel C. Melchior, American Chemical
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