

Chapter-4

Reaction Stoichiometry

- Stoichiometry is the study of the quantitative relationships between amounts of reactants used and amounts of products formed by a chemical reaction.
 - Chemical equations
 - The mole and molar mass
 - Chemical formulas
 - Mass relationships in equations
 - Limiting reactant

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4.1. Interpreting Stoichiometric Coefficients

$4\text{Fe(s)} + 3\text{O}_2\text{(g)}$	\rightarrow	$2\text{Fe}_2\text{O}_3\text{(s)}$
Iron + oxygen	\rightarrow	Iron (III) oxide
4 atoms of Fe + 3 molecules of O_2	\rightarrow	2 formula units of Fe_2O_3
4 mol of Fe + 3 mol of O_2	\rightarrow	2 mol Fe_2O_3
223.4 g of Fe + 96 g of O_2	\rightarrow	319.4 g of Fe_2O_3
319.4 g reactants	\rightarrow	319.4 g products

- The chemical equation does not directly tell you anything about the masses of reactants and products. However, by converting the known mole quantities to mass.
 - Moles are converted to mass by multiplying their molar mass

$$4 \text{ mol Fe} \times \frac{55.85 \text{ g Fe}}{1 \text{ mol Fe}} = 223.4 \text{ g Fe}$$

Reaction Stoichiometry Lab Answers

**James M. Postma, Julian L. Roberts, J.
Leland Hollenberg**

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