

Vol. of 1 mol of ideal gas at STP/NTP =  $22.4 \text{ L}$

उत्स नान एरिया  
- 3 same sabse kam significant figure hoga pt. ke bad utne hi hoge ans me.

\* Significant fig  
• 0 to the left of the 1st non 0 digit are sig. not sig.  
eg  $0.0012 \Rightarrow 2$   
• Terminal 0 are not sig. if there is no decimal pt.  
eg  $1000 \Rightarrow 1$  fig.  
• 0 at end is sig. if there is decimal pt.  
• Exponential term are not considered sig.

अंश - 3/131  
 $a \times b = c$  } - add/subst  
is least amount  
3 same sabse kam fig hoga vo sig. fig ans me hoga

\* LAW OF MASS CONSV  
• In a react<sup>n</sup>  $\rightarrow$  actual mass of reactant = mass of prod.  
Rakme zitna use hua.  
• Except<sup>n</sup>  $\rightarrow$  Nuclear Rxn ( $E = mc^2$ )  
• PppWed<sup>n</sup>  $\rightarrow$  Balanced Rxn

# MOLE CONCEPT

\* LAW OF CONST PROPORT<sup>n</sup>  
• Elements in a comp are in fixed ratio by mass.  
eg  $\text{CO}_2$   
 $12 \quad 32 \text{g} \Rightarrow 3:8$   
Not valid for isotope.  
 $\% \text{ of C} = \frac{3}{11} \times 100$   
Q. cal. the mass of C req to form 100g of  $\text{CO}_2$ .  
 $\Rightarrow \text{C} + 2\text{O} \rightarrow \text{CO}_2$   
 $12 \text{g} \quad 32 \text{g} \quad 44 \text{g}$   
 $44 \text{g CO}_2$  ke lie  $12 \text{g C}$   
 $\therefore 100 \text{g ke lie} = \frac{12}{44} \times 100 = 30 \text{g}$

\* Law of Multiple Prop<sup>n</sup>  
• elements should be same or same fixed mass  
eg.  $\text{CO}$   $\text{CO}_2$  1:2 ratio  
Other eg.  $\text{N}_2\text{O}$ ,  $\text{N}_2\text{O}_3$ ,  $\text{N}_2\text{O}_5$   
except<sup>n</sup>  $\rightarrow \text{N}_2\text{O}_3$   
\* Avogadro's law  
• Same (P, T, R, V) of molecules of gases some (mass eg. no. of molecules)  
 $\text{Xe}_2, \text{Xe}_3, \text{Xe}_4$   
 $2 : 3 : 4$  molecules  
 $2 : 3 : 4$  atoms (different)

\* Gay Lussac's law  
CONDINS: +T, P some  
• at least 2 gases.  
 $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$   
 $1 : 1 : 2$   
• Vol.  $\text{H}_2$  lie

\* DALTON'S ATOMIC THEORY  
• Atoms  $\rightarrow$  Neither created nor destroyed

\* MOLE CONCEPT  
Atomic No. (P)  $\rightarrow$  mass no. (P+N)  
 $\frac{A}{Z} \times (P+N)$   
• Mass of Proton =  $1.67 \times 10^{-27} \text{ kg}$   
• Neutron =  $1.67 \times 10^{-27} \text{ kg}$   
• Electron =  $9.1 \times 10^{-31} \text{ kg}$   
•  $m_p \approx 1837 m_e$

\* Atomic mass unit (amu)  
 $1 \text{ amu} = \frac{1}{12}$  Mass of  $^{12}\text{C}$   
 $1 \text{ amu} = 1.67 \times 10^{-27} \text{ kg}$   
eg Mass of  $^{12}\text{C} = 12 \text{ amu}$   
\* Relative atomic mass  
• Unitless  
• eg. mass. A.M. of C = 12

\* Molecular Mass  
•  $\text{H}_2\text{SO}_4 = 98 \text{ amu}$   
 $\rightarrow$  Mass of 1 molecule of  $\text{H}_2\text{SO}_4$   
\* FORMULA MASS: for ionic comp.

\* MOLE  
Amount of substance.  
• 1 mole = No particles ( $6.022 \times 10^{23}$ )  
 $1 \text{ amu} \leq 1.67 \times 10^{-27} \text{ kg}$   
 $\text{Na} = 6.022 \times 10^{23}$   
 $1 \text{ amu} \text{ or } \text{Na} \leq 1 \text{ g}$

# Mole Concept Theory Notes Pdfslibforme

**Wolfgang Guggemos**



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