

## Ideal gas law - problems and solutions

1. Ideal gases in a closed container initially have volume  $V$  and temperature  $T$ . The final temperature is  $5/4T$  and the final pressure is  $2P$ . What is the final volume of gases.

Known :

Initial volume ( $V_1$ ) =  $V$

Initial temperature ( $T_1$ ) =  $T$

Final temperature ( $T_2$ ) =  $5/4 T$

Initial pressure ( $P_1$ ) =  $P$

Final pressure ( $P_2$ ) =  $2P$

Wanted : Final volume ( $V_2$ )

Solution :

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$\frac{P V}{T} = \frac{(2P) V_2}{5/4 T}$$

$$\frac{V}{1} = \frac{2 V_2}{5/4}$$

$$5/4 V = 2 V_2$$

$$V_2 = \frac{5V}{4} \cdot \frac{2}{1}$$

$$V_2 = \frac{5V}{4} \times \frac{1}{2}$$

$$V_2 = \frac{5V}{8}$$

2. Determine the volume of 2.00 moles of gases (ideal gas) at STP. STP = Standard Temperature and Pressure.

Known :

Moles of gas ( $n$ ) = 2 moles

Standard temperature ( $T$ ) =  $0^\circ\text{C} = 0 + 273 = 273$  Kelvin

Standard pressure ( $P$ ) = 1 atm =  $1.013 \times 10^5$  Pa

Universal gas constant ( $R$ ) = 8.315 Joule/mole.Kelvin

Wanted : Volume of gases ( $V$ )

Solution :

Read more :

<https://physics.gurumuda.net/ideal-gas-law-problems-and-solutions.htm>

# Gas Law Problems With Solutions

**IM Harris**



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### **Table of Contents Gas Law Problems With Solutions**

1. Understanding the eBook Gas Law Problems With Solutions
  - The Rise of Digital Reading Gas Law Problems With Solutions
  - Advantages of eBooks Over Traditional Books
2. Identifying Gas Law Problems With Solutions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gas Law Problems With Solutions
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Law Problems With Solutions
  - Personalized Recommendations
  - Gas Law Problems With Solutions User Reviews and Ratings

- Gas Law Problems With Solutions and Bestseller Lists
- 5. Accessing Gas Law Problems With Solutions Free and Paid eBooks
  - Gas Law Problems With Solutions Public Domain eBooks
  - Gas Law Problems With Solutions eBook Subscription Services
  - Gas Law Problems With Solutions Budget-Friendly Options
- 6. Navigating Gas Law Problems With Solutions eBook Formats
  - ePub, PDF, MOBI, and More
  - Gas Law Problems With Solutions Compatibility with Devices
  - Gas Law Problems With Solutions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gas Law Problems With Solutions
  - Highlighting and Note-Taking Gas Law Problems With Solutions
  - Interactive Elements Gas Law Problems With Solutions
- 8. Staying Engaged with Gas Law Problems With Solutions
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Gas Law Problems With Solutions
- 9. Balancing eBooks and Physical Books Gas Law Problems With Solutions
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gas Law Problems With Solutions
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Gas Law Problems With Solutions
  - Setting Reading Goals Gas Law Problems With Solutions
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gas Law Problems With Solutions
  - Fact-Checking eBook Content of Gas Law Problems With Solutions
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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