

Study Card: Quantum Numbers

- In Quantum Mechanics, the electronic structure of atoms is described by a wave function.
- Valid solutions to the wave function yield a unique set of four numerical values (known as **quantum numbers**) to describe an electron and its region of space (i.e., **orbital**).
- Each quantum number has a name and a set of restrictions which govern the allowed numerical values that an electron can have.

#1 Principal Quantum Number (n)

Allowed values: $n = 1, 2, 3, 4, \dots$

- Describes the energy level and the size of the orbitals.
- Determines the # of orbital types in that energy level (n), the total # of orbitals (n^2), and the # of e's ($2n^2$)

#2 Azimuthal Quantum Number (l)

Allowed values: $0, 1, 2, \dots, n-1$

e.g., if $n = 3$, allowed values are $\{0, 1, 2\}$

- Describes the type of orbital
- Determines # of orbitals of that type ($2l + 1$)
- $l = 0 \rightarrow$ s orbital (only 1)
- $l = 1 \rightarrow$ p orbital (three)
- $l = 2 \rightarrow$ d orbital (five)
- $l = 3 \rightarrow$ f orbital (seven)

#3 Magnetic Quantum Number (m_l)

Allowed values: $-l, -l + 1, \dots, l-1, l$

e.g., if $l = 2$, allowed values are $\{-2, -1, 0, 1, 2\}$

- Describes spatial orientation of the orbital
- The # of m_l values = # of orbitals of that type

#4 Electron Spin Quantum Number (m_s)

Allowed values: $-\frac{1}{2}, +\frac{1}{2}$

- Describes the electron's spin
- Spin is described as up or down

n or Electron Shell	# of Orbital Types	# of Orbitals	Max. # of Electrons
1 ... 1 st electron shell	1	1	2
2 ... 2 nd electron shell	2	4	8
3 ... 3 rd electron shell	3	9	18
4 ... 4 th electron shell	4	16	32
n ... n^{th} electron shell	n	n^2	$2 \cdot n^2$

n	# of Orbital Types	l	Orbital Type	# of These Orbitals
1	1 (s only)	0	1s	1
2	2 (s and p)	0	2s	1
		1	2p	3
3	3 (s, p, and d)	0	3s	1
		1	3p	3
		2	3d	5
4	4 (s, p, d, and f)	0	4s	1
		1	4p	3
		2	4d	5
		3	4f	7

n	Possible l values	l	Orbital Type	# of These Orbitals	Possible m_l values
1	{0}	0	1s	1	{0}
		0	2s	1	{0}
2	{0,1}	1	2p	3	{-1,0,+1}
		0	3s	1	{0}
		1	3p	3	{-1,0,+1}
3	{0,1,2}	2	3d	5	{-2,-1,0,+1,+2}
		0	4s	1	{0}
		1	4p	3	{-1,0,+1}
		2	4d	5	{-2,-1,0,+1,+2}
4	{0,1,2,3}	3	4f	7	{-3,-2,-1,0,+1,+2,+3}

Deciphering Quantum Numbers

Numbers ordered as $\{n, l, m_l, m_s\}$

$\{2,0,0,+\frac{1}{2}\} \rightarrow$ e⁻ in 2s orbital

$\{3,0,0,+\frac{1}{2}\} \rightarrow$ e⁻ in 3s orbital

$\{2,1,-1,-\frac{1}{2}\} \rightarrow$ e⁻ in 2p orbital

$\{3,1,1,+\frac{1}{2}\} \rightarrow$ e⁻ in 3p orbital

$\{3,2,1,-\frac{1}{2}\} \rightarrow$ e⁻ in 3d orbital

$\{4,2,-2,+\frac{1}{2}\} \rightarrow$ e⁻ in 4d orbital

$\{4,3,2,+\frac{1}{2}\} \rightarrow$ e⁻ in 4f orbital

Red #s above are one of many allowed values for that quantum number.

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