

Stoichiometry & the mole notes

the mole concept

the mole (mol) is the SI unit for measuring the amount of a substance.

avogadro's number:

$$1 \text{ mol} = 6.022 \times 10^{23} \text{ particles}$$

(particles can be atoms, molecules, or formula units)

molar mass:

the mass of one mole of a substance in grams.

found by adding the atomic masses from the periodic table

example

$$\text{H}_2\text{O} (2 \cdot 1.008 \text{ g/mol}) + (1 \cdot 16.00 \text{ g/mol}) = 18.016 \text{ g/mol}$$

Example Problems

1. how many atoms are in 3.5 moles of Fe?

$$3.5 \text{ mol Fe} \cdot \frac{6.022 \cdot 10^{23} \text{ atoms}}{1 \text{ mol}} = 2.11 \cdot 10^{24} \text{ atoms}$$

2. what is the mass of 2.00 moles of NaCl?

find molar mass:

$$\text{Na} : 22.99 \text{ g/mol}$$

$$\text{Cl} : 35.45 \text{ g/mol}$$

$$\text{Total} : 58.44 \text{ g/mol}$$

convert

$$2.00 \text{ mol} \cdot \frac{58.44 \text{ g}}{1 \text{ mol}} = 116.88 \text{ g}$$



Conversions Involving the mole

conversion

particles \leftrightarrow moles

moles \leftrightarrow mass

moles \leftrightarrow volume (L) (at STP)

factor to use

$$1 \text{ mol} = 6.022 \cdot 10^{23} \text{ particles}$$

$$1 \text{ mol} = \text{molar mass (g)}$$

$$1 \text{ mol} = 22.4 \text{ L of gas}$$

Stoichiometry: the study of quantitative relationships in chemical reactions

Balanced chemical equations show the mole ratio of reactants & products

mole ratios come from coefficients in a balanced equation.

Stoichiometry steps:

1. Balance the equation

2. convert given quantity into moles (if necessary)

3. Use mole ratio from balanced equation.

4. Convert moles back to desired unit (grams, liters, etc.)

Example:



question: how many grams of H_2O are produced from 5.00 moles of

O_2 ? Step 1: Use mole ratio

$$5.00 \text{ mol O}_2 \cdot \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol O}_2} = 10.0 \text{ mol H}_2\text{O}$$

Step 2: Convert to grams

$$10.0 \text{ mol H}_2\text{O} \cdot \frac{18.016 \text{ g}}{1 \text{ mol}} = 180.16 \text{ g H}_2\text{O}$$

Limiting Reactant & Percent Yield

limiting reactant (LR): the reactant that runs out first, limiting product formation

excess reactant: the reactant that remains after the reaction is complete

Steps to identify LR

1. Convert reactants to moles

2. Use mole ratios to do

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JS Bruner



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Chemistry of Pyrotechnics Chris Mocella, John A. Conkling, 2019-01-15 This book provides chemists with technical insight on pyrotechnics and explosives It emphasizes basic chemical principles and practical hands on knowledge in the preparation of energetic materials It examines the interactions between and adaptations of pyrotechnics to changing technology in areas such as obscuration science and low signature flame emission The updated third edition discusses chemical and pyrotechnic principles components of high energy materials elements of ignition propagation and sensitivity It offers heat compositions including ignition mixes delays thermites and propellants and investigates the production of smoke and sound as well as light and color *Lecture-notes on Theoretical Chemistry* Ferdinand Gerhard Wiechmann, 1893 [High School Chemistry](#)

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