

9

CHEMICAL NAMES AND FORMULAS

SECTION 9.1 NAMING IONS (pages 253–258)

This section explains the use of the periodic table to determine the charge of an ion. It also defines polyatomic ion and gives the names and formulas for the most common polyatomic ions.

► Monatomic Ions (pages 253–256)

1. What are monatomic ions?

Monatomic ions are ions consisting of only one atom.

2. How is the ionic charge of a Group 1A, 2A, or 3A ion determined?

The ionic charge is numerically equal to the group number.

3. How is the ionic charge of a Group 5A, 6A, or 7A ion determined?

The charge of an ion in Groups 5A, 6A, or 7A is determined by subtracting 8 from the group number.

4. Circle the letter of the type of element that often has more than one common ionic charge.

- a. alkali metal
- b. alkaline earth metal
- c. transition metal
- d. nonmetal

5. The _____ **Stock system** _____ of naming transition metal cations uses a Roman numeral in parentheses to indicate the numeric value of the ionic charge.

6. An older naming system uses the suffix **-ous** to name the cation with the _____ **lesser** _____ charge, and the suffix **-ic** to name the cation with the _____ **greater** _____ charge.

7. What is a major advantage of the Stock system over the old naming system?

The Stock system gives the actual charge of the ion.

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WJ Hussar



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Table of Contents Chemistry Chapter 9 Chemical Names Formulas Answers

1. Understanding the eBook Chemistry Chapter 9 Chemical Names Formulas Answers
 - The Rise of Digital Reading Chemistry Chapter 9 Chemical Names Formulas Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 9 Chemical Names Formulas Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 9 Chemical Names Formulas Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 9 Chemical Names Formulas Answers
 - Personalized Recommendations
 - Chemistry Chapter 9 Chemical Names Formulas Answers User Reviews and Ratings
 - Chemistry Chapter 9 Chemical Names Formulas Answers and Bestseller Lists
5. Accessing Chemistry Chapter 9 Chemical Names Formulas Answers Free and Paid eBooks
 - Chemistry Chapter 9 Chemical Names Formulas Answers Public Domain eBooks
 - Chemistry Chapter 9 Chemical Names Formulas Answers eBook Subscription Services
 - Chemistry Chapter 9 Chemical Names Formulas Answers Budget-Friendly Options

6. Navigating Chemistry Chapter 9 Chemical Names Formulas Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 9 Chemical Names Formulas Answers Compatibility with Devices
 - Chemistry Chapter 9 Chemical Names Formulas Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 9 Chemical Names Formulas Answers
 - Highlighting and Note-Taking Chemistry Chapter 9 Chemical Names Formulas Answers
 - Interactive Elements Chemistry Chapter 9 Chemical Names Formulas Answers
8. Staying Engaged with Chemistry Chapter 9 Chemical Names Formulas Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 9 Chemical Names Formulas Answers
9. Balancing eBooks and Physical Books Chemistry Chapter 9 Chemical Names Formulas Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 9 Chemical Names Formulas Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 9 Chemical Names Formulas Answers
 - Setting Reading Goals Chemistry Chapter 9 Chemical Names Formulas Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter 9 Chemical Names Formulas Answers
 - Fact-Checking eBook Content of Chemistry Chapter 9 Chemical Names Formulas Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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